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 $2\text{ nacl} + \text{f}_2 \rightarrow 2\text{ naf} + \text{cl}_2$
 $4\text{ h}_2 + \text{o}_2 \rightarrow 2\text{ h}_2\text{o}$
 $5\text{ pb} + \text{oh}_2 + 2\text{ hcl} \rightarrow 2\text{ h}_2\text{o} + 1\text{ pbcl}_2 + 2\text{ albr}_3 + 3\text{ k}_2\text{so}_4 + 6\text{ kbr} + 1\text{ al}_2\text{so}_4 + 3\text{ h}_2 + 4\text{ ch}_4$

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 $3\text{ hbr} + \text{khco}_3 + \text{h}_2\text{o} \rightarrow \text{kbr} + \text{co}_2 + \text{gabr} + \text{nabr} + \text{sno}_4$
 $\text{naso}_3 + \text{gaz} + \text{so}_3 \rightarrow 3\text{ nf}_3 + 2\text{ snf}_2 + \text{n}_2\text{o}_3$ using the equation from problem 2 above answer the following

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 $2\text{ c}_6\text{h}_{10} + 17\text{ o}_2 \rightarrow 12\text{ co}_2 + 10\text{ h}_2\text{o}$
 $3\text{ hbr} + 1\text{ khco}_3 + 1\text{ h}_2\text{o} \rightarrow 1\text{ kbr} + \text{co}_2 + \text{gabr}_3 + 3\text{ na}_2\text{so}_3 + 1\text{ ga}_2\text{so}_3 + 3\text{ so}_3 + 6\text{ nabr} + 3\text{ sno}_2 + 2\text{ nf}_3 + 3\text{ snf}_2 + 1\text{ n}_2\text{o}_3$ using the following equation
 $2\text{ naoh} + \text{h}_2\text{so}_4 \rightarrow 2\text{ h}_2\text{o} + \text{na}_2\text{so}_4$

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