



Method Implementation Document for EN 13284-1

BS EN 13284-1:2002

**Stationary source emissions – Determination of low range mass
concentration of dust – Part 1: Manual gravimetric method**

Environment Agency

Version 2.4

December 2011



Method Implementation Document For En 13284 1

SJ Ball



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) + 5 \text{ O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g}) + \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) + \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) + \Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) + \Delta h = 1560 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) + \Delta h = 285 \text{ kJ}$

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 $\text{C}_2\text{N}_2(\text{g}) + 2 \text{ O}_2(\text{g}) + \Delta h = 2 \text{ NO}_2(\text{g}) + \text{N}_2\text{O}_5(\text{g})$
given the following information

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hess s law is written as $\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$ or $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

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to use hess s law two principles must be understood
one if an equation is reversed the sign of the Δh value is also reversed
two if an equation is multiplied by a

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{ O}_2(\text{g}) + \Delta h = \text{MgO}(\text{s})$
we used

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are
cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq})$

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c_2h_4

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web thus for reaction 9 2 in which 1 mole of h_3o^+ and 1 mole of oh^- combine to form 2 moles of h_2o h 55.8 kj as shown note that the enthalpy change $h_{products} - h_{reactants}$ is positive if heat is absorbed that is if $h_{products}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in

solution eventually allowing an individual to prove or disprove Hess's Law. Introduction: Thermodynamics is the study of energy exchanged between the system and the surroundings.

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web 3 12 Hess's Law in this experiment you will measure the amount of heat released in these three related exothermic reactions:

1. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
2. $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$
3. $\text{H}^+ \text{(aq)} + \text{OH}^- \text{(aq)} \rightarrow \text{H}_2\text{O (l)}$