



$$\Delta H_{298}^{\circ} = -68.73 \text{ kcal } (-287.57 \text{ kJ})$$

Atomic Weight of **B** = 10.82 grams.

Atomic Weight of **O** = 16.0 grams.

**Required:**

- (a) Standard heat of formation  $\Delta H_{298}^{\circ}$  of  $\text{B}_2\text{O}_3$  in terms of per mole of  $\text{B}_2\text{O}_3$  =?
- (b) Standard heat of formation  $\Delta H_{298}^{\circ}$  of  $\text{B}_2\text{O}_3$  in terms of per gram of  $\text{B}_2\text{O}_3$  =?

**Solution:**

Now, we will multiply above given equations with real numbers to make them equal and then we will add all equations with each other. It will yield the final equation which will lead us to the final solution of this problem.

**1. Multiply eq (i) with 2.**



# Metallurgical Thermodynamics Problems And Solution

**A Gutmann**



## **Metallurgical Thermodynamics Problems And Solution:**

## Unveiling the Magic of Words: A Report on "**Metallurgical Thermodynamics Problems And Solution**"

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