

Solutions Manual to Accompany
**Basic Principles and Calculations
in Chemical Engineering**

Seventh Edition

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PRENTICE HALL
Professional Technical Reference
Upper Saddle River, New Jersey 07458

TYPICAL EXAMS FOR A ONE SEMESTER COURSE
(Scheduled in the evening to avoid time constraints on students)

Exam No. 1
(Open Book, 1 1/2 hours)

PROBLEM 1 (25%)

Hydrogen can be separated from natural gas by diffusion through a metal tube. The rate of separation is given by

$$N = 2 \pi D p R$$

where
 N = rate of transport of H_2 from the tube, g moles/decade (of length of tube)
 D = diffusion coefficient
 p = molar density of H_2 , g moles/cm³
 R = log mass radius of tube, $r_2 - r_1$ is $\left(\frac{r_2}{r_1}\right)$, with r in cm.

What are the units of D ?

PROBLEM 2 (25%)

A pellet of brass weighing 10 grams is dropped from a lift truck from a height of 10 feet. The maximum velocity that pellet attains before hitting the ground is 8 ft/sec. How much kinetic energy does the pellet have in (Joules) at this velocity?

PROBLEM 3 (25%)

The specific gravity of a fuel oil is 0.82. What is the density of the oil in lb/ft³? Show all units.

PROBLEM 4 (25%)

Sulfur trioxide (SO_3) can be absorbed in sulfuric acid solution to form more concentrated sulfuric acid. If the gas to be absorbed contains 50% SO_3 , 41% N_2 , 7% SO_2 , and 1% O_2 , how many parts per million of O_2 are there in the gas? (a) What is the composition of the gas on a % free basis?

PROBLEM 5 (25%)

You have 100 kilograms of gas of the following composition:

CH_4	30%
H_2	20%
N_2	50%

What is the average molecular weight of this gas?

PROBLEM 6 (25%)

If the heat capacity of a substance is 3.22 J/g(°C) and its molecular weight is 75.4, what is its heat capacity in

(a) Btu(°F)
(b) J/kg(°K)
(c) J/gmol(°K)

PROBLEM 7 (25%)

A rock containing 100% $BaSO_4$ is heated with coke (94% C, 6% ash) and the composition of the product is $BaSO_4$ (11.5%), BaS (72.9%), C(13.9%), ash(2.2%). The reaction is



Calculate the percent excess reactant, and the degree of completion of the reaction.

PROBLEM 8 (25%)

A gas cylinder to which is attached an Bourdon gage appears to be at a pressure of 21.38 in. Hg at 70°F. The barometer reads 141.8 inHg. A student claims that the pressure in the tank is 1.3 psia, but another student points out that this is impossible - the pressure is really 28.2 psia. Can 1.3 psia be correct? Explain and show calculations to back up your explanation.

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Table of Contents Solution Basic Principles Himmelblau Ed 6

1. Understanding the eBook Solution Basic Principles Himmelblau Ed 6
 - The Rise of Digital Reading Solution Basic Principles Himmelblau Ed 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Basic Principles Himmelblau Ed 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Basic Principles Himmelblau Ed 6
 - User-Friendly Interface

4. Exploring eBook Recommendations from Solution Basic Principles Himmelblau Ed 6
 - Personalized Recommendations
 - Solution Basic Principles Himmelblau Ed 6 User Reviews and Ratings
 - Solution Basic Principles Himmelblau Ed 6 and Bestseller Lists
5. Accessing Solution Basic Principles Himmelblau Ed 6 Free and Paid eBooks
 - Solution Basic Principles Himmelblau Ed 6 Public Domain eBooks
 - Solution Basic Principles Himmelblau Ed 6 eBook Subscription Services
 - Solution Basic Principles Himmelblau Ed 6 Budget-Friendly Options
6. Navigating Solution Basic Principles Himmelblau Ed 6 eBook Formats
 - ePub, PDF, MOBI, and More
 - Solution Basic Principles Himmelblau Ed 6 Compatibility with Devices
 - Solution Basic Principles Himmelblau Ed 6 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Basic Principles Himmelblau Ed 6
 - Highlighting and Note-Taking Solution Basic Principles Himmelblau Ed 6
 - Interactive Elements Solution Basic Principles Himmelblau Ed 6
8. Staying Engaged with Solution Basic Principles Himmelblau Ed 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Basic Principles Himmelblau Ed 6
9. Balancing eBooks and Physical Books Solution Basic Principles Himmelblau Ed 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Basic Principles Himmelblau Ed 6
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Basic Principles Himmelblau Ed 6
 - Setting Reading Goals Solution Basic Principles Himmelblau Ed 6
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Solution Basic Principles Himmelblau Ed 6
 - Fact-Checking eBook Content of Solution Basic Principles Himmelblau Ed 6
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web this article will outline how to start a bottled water business and the bottled water business plan pdf word excel market
research the most important aspect you will look at here is finding out about existing players to draw insights from them

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web written by dave lavinsky bottled water business plan over the past 20 years we have helped over 1 000 entrepreneurs
and business owners create business plans to start and grow their bottled water production companies on this page we will
first give you some background information with regards to the importance of business planning