

Gravimetric Analysis Example – try it!

- An excess of aqueous sodium phosphate is added to a 50.0 mL sample of copper (II) chloride, with unknown concentration. A precipitate forms, as shown in the molecular equation below:



A student purifies, dries, and weighs the solid precipitate. She weighs the solid and finds she has collected 18.0 grams of $\text{Cu}_3(\text{PO}_4)_2$. Based on this data, determine the molarity of the CuCl_2 solution.

$$18.0 \text{ g } \cancel{\text{Cu}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol } \cancel{\text{Cu}_3(\text{PO}_4)_2}}{380.4 \text{ g } \cancel{\text{Cu}_3(\text{PO}_4)_2}} \times \frac{3 \text{ mol } \text{CuCl}_2}{1 \cancel{\text{Cu}_3(\text{PO}_4)_2}} = 0.142 \text{ mol } \text{CuCl}_2$$

$$M = \frac{\text{mol}}{\text{L}} = \frac{0.142 \text{ mol}}{0.0500 \text{ L}} = 2.84$$

Stoichiometry And Gravimetric Analysis Lab Answers

David Callaghan



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