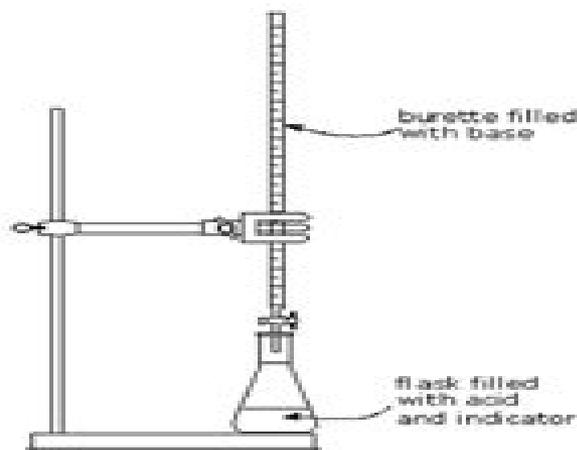


# Titration Lab

**Problem:** To determine the concentration of the unknown acid or base.

**Set-up:**



**Procedure:**

1. Place 10 mL of HCl in the Erlenmeyer flask.
2. Place 5 drops of phenolphthalein in the Erlenmeyer flask with the HCl.
3. Fill the burette with NaOH using a funnel.
4. Place the burette in the retort stand to look like the set up in the diagram.
5. Place a piece of white paper underneath the Erlenmeyer flask.
6. Write down the initial volume of NaOH on the burette.
7. Slowly (drop by drop) add the NaOH to the Erlenmeyer flask, while swirling the liquid, until the solution stays a very light pink colour.
8. Write down the final volume of the NaOH on the burette.
9. Rinse out the Erlenmeyer flask and repeat the experiment at least 3 times.

**Observations (4):**

	Trial 1	Trial 2	Trial 3	Trial 4
Initial (I) Volume of NaOH (mL)	0.00 mL	11.1 mL	22.5 mL	33.8 mL
Final (F) Volume of NaOH (mL)	11.1 mL	22.5 mL	33.8 mL	45.3 mL
Total Volume of NaOH used (mL) (I-F)*	11.1 mL	11.4 mL	11.3 mL	11.5

\* the burette is labelled as full at 50 mL and goes down from there, so you will subtract the final value from the initial to get the total volume used.

# Titration Lab Answers

**Glencoe**



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