

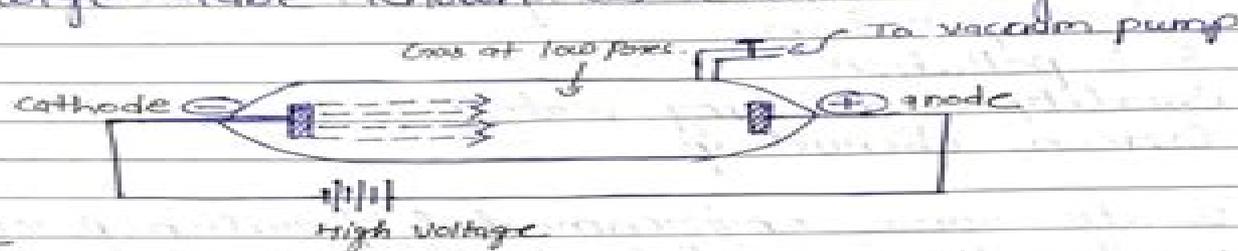
CHAPTER - 2

STRUCTURE OF ATOM

Discovery of electron (Cathode Rays)

Cathode rays \rightarrow William Crookes (1879)
Electron (e^-) \rightarrow J. J. Thomson (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



properties

- 1 Cathode rays travel in a straight line (shadow)
- 2 Cathode rays carry -ve charge (electrostatic field deflection towards anode)
- 3 Cathode rays are deflected by magnetic field (rotate paddle wheel)
- 4 They are made up of material particles
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (when they strike metal)
- 7 They also affect photographic plate.

Discovery of proton (Anode / Canal Rays)

Anode rays \rightarrow Goldstein (1886)
Proton (H^+) \rightarrow Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

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Wolfgang Guggemos



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1. Understanding the eBook Class 11th Notes Chemistry Chapter Wise
 - The Rise of Digital Reading Class 11th Notes Chemistry Chapter Wise
 - Advantages of eBooks Over Traditional Books
2. Identifying Class 11th Notes Chemistry Chapter Wise
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Class 11th Notes Chemistry Chapter Wise
 - User-Friendly Interface
4. Exploring eBook Recommendations from Class 11th Notes Chemistry Chapter Wise
 - Personalized Recommendations
 - Class 11th Notes Chemistry Chapter Wise User Reviews and Ratings
 - Class 11th Notes Chemistry Chapter Wise and Bestseller Lists
5. Accessing Class 11th Notes Chemistry Chapter Wise Free and Paid eBooks
 - Class 11th Notes Chemistry Chapter Wise Public Domain eBooks
 - Class 11th Notes Chemistry Chapter Wise eBook Subscription Services
 - Class 11th Notes Chemistry Chapter Wise Budget-Friendly Options
6. Navigating Class 11th Notes Chemistry Chapter Wise eBook Formats

- ePub, PDF, MOBI, and More
 - Class 11th Notes Chemistry Chapter Wise Compatibility with Devices
 - Class 11th Notes Chemistry Chapter Wise Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Class 11th Notes Chemistry Chapter Wise
 - Highlighting and Note-Taking Class 11th Notes Chemistry Chapter Wise
 - Interactive Elements Class 11th Notes Chemistry Chapter Wise
 8. Staying Engaged with Class 11th Notes Chemistry Chapter Wise
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Class 11th Notes Chemistry Chapter Wise
 9. Balancing eBooks and Physical Books Class 11th Notes Chemistry Chapter Wise
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Class 11th Notes Chemistry Chapter Wise
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Class 11th Notes Chemistry Chapter Wise
 - Setting Reading Goals Class 11th Notes Chemistry Chapter Wise
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Class 11th Notes Chemistry Chapter Wise
 - Fact-Checking eBook Content of Class 11th Notes Chemistry Chapter Wise
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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