



Electronics Repair

Electronics Repair Guide Step By Step

JL Elias



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alcohols ethanol 2 methylpropan 2 ol butan 2 ol but 2 en 1 ol ch₃ch₂ch₂ch₂oh ii ch₃ ch₃ch₂ch₂oh iii ch₃ch₃ch₂ch₂oh iv ch₃ ch₃ch₂ch₂oh ch₃ 6 aldehydes and ketones draw the structures for the following compounds ethanal butanone also known as butan 2 one but the 2 is unnecessary and is often left out 2 methylpropanal

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question 1 basic hydrocarbon nomenclature part a enter the number of carbons indicated by each of the following root names in the designated answer box part b the line formula for a branched alkane is shown below part c the line formula for another branched alkane is

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jan 21 2020 the point of attachment is number one and any other smaller groups are named as substituent groups on the larger group this numbering is independent of the numbering of the parent chain try to name the following compounds

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naming organic compounds 1 1 alkanes a i you shouldn't really have got this one wrong unless you made a careless mistake like putting too many or too few hydrogens on the middle carbon if you got it wrong re read the bit about naming alkanes and then look at these questions again ii

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dec 16 2021 so the complete name is 5 bromo 7 chloro 6 hydroxy 2 2 5 trimethyl 7 octen 4 one 4 it is not difficult to find the parent structure for this compound which is a cyclic alcohol so the last name is cyclopropanol the naming of the substituent with the benzene ring is

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solutions 1 aromatic compound iodobenzene 2 alkane 3 methylhexane 3 ketone 4 alkane alkyl halide 3 chloro 4 ethyl 2 4 dimethyloctane 5 aromatic compound o diethylbenzene or ortho diethylbenzene 6 alkene 3 methylpentene 7 ester 8 aromatic compound p bromotoluene or para bromotoluene 9 alkyne

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jan 23 2023 jim clark truro school in cornwall this page explains how to write the formula for an organic compound given its name and vice versa it covers alkanes cycloalkanes alkenes simple compounds containing halogens alcohols aldehydes and ketones

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feb 21 2017 give the complete iupac name for each of the following compounds a b c give the complete iupac name and the common name for each of the following compounds a b c a 1 ethoxybutane butyl ethyl ether b 1 ethoxypropane ethyl propyl ether c 1 methoxypropane methyl propyl ether

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this is a set of practice problems on naming organic compounds the examples cover the nomenclature of alkanes bicyclic compounds alkenes alkynes alcohols alkyl halides aromatic compounds aldehydes and ketones amines ethers and carboxylic acid derivatives such as nitriles esters and amides

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nomenclature of organic compounds chemistry questions with solutions q1 the iupac name of CH_3CHO is ethanal ethanol acetaldehyde acetone answer a the iupac name of CH_3CHO is ethanal q2 the iupac name of $\text{CH}_3\text{OCH}_2\text{CH}_2\text{CH}_3$ is methoxy ethane ethoxy ethane methyl ethane ethyl ethane

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1 carbon methane 2 carbons ethane 3 carbons propane 4 carbons butane 5 carbons pentane 6 carbons hexane 7 carbons

heptane 8 carbons octane 9 carbons nonane 10 carbons decane while many of these names share a greek root with

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web a nuclear reaction is a reaction that affects the nucleus of an atom one type of a nuclear reaction is radioactive decay a reaction in which a nucleus spontaneously disintegrates into a slightly lighter nucleus accompanied by the emission of particles energy or both an example is shown below in which the nucleus of a polonium atom

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web a nuclear reaction is a process that occurs when a nuclear particle nucleon or nucleus gets into close contact with another most of the known nuclear reactions are produced by exposing different materials to a beam of accelerated nuclear particles

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web aug 13 2020 in an atomic bomb or fission bomb the nuclear chain reaction is designed to be uncontrolled releasing huge amounts of energy in a short amount of time a critical mass of fissionable plutonium is contained within the bomb but not at a sufficient density

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web the reaction in our example above would be written as ${}_{6}^{14}\text{Li} \rightarrow {}_{4}^{10}\text{Be} + {}_{2}^{4}\text{He}$ balancing a radioactive decay equation in balancing a nuclear equation it is important to remember that the sum of all the mass numbers and atomic numbers given on the upper left and lower left side of the element symbol respectively must be equal for both sides of the

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capture reactions free neutrons and protons fused to deuterium d or $2h$ with the excess energy emitted as a 2.2 mev gamma ray $n + p \rightarrow d + \gamma$

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