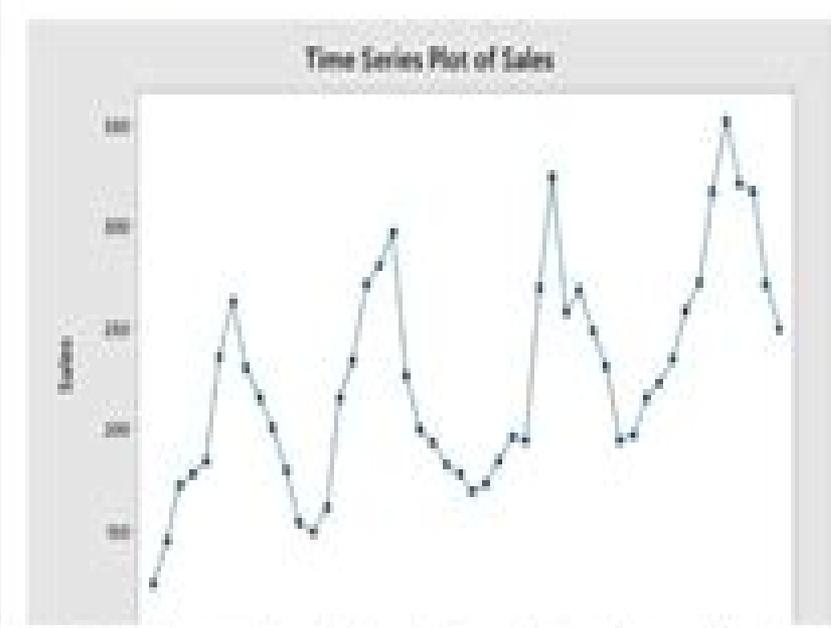


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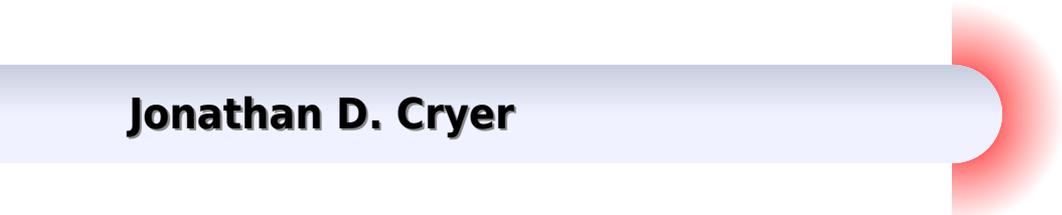


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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom s oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom s oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{CO}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{C}$
oxidation $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$ reduction $\text{Fe}_2\text{O}_3 + 2\text{e}^- \rightarrow 2\text{Fe} + 3\text{O}^{2-}$ $24 \text{ Ag} + \text{Ni}^{3+} \rightarrow 3\text{Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu} + \text{Pb}^{2+} \rightarrow \text{Cu}^{2+} + \text{Pb}$ oxidation reduction 26 O_2

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent $\text{a} + \text{i}_2\text{o}_5 \rightarrow 5\text{CO}_2 + \text{i}_2 + \text{O}_2$ $5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced $\text{Ox Ag} + 2\text{Li} \rightarrow \text{Li}_2\text{Ag} + \text{Li}^0$ oxidized red ag S^0 to S^{2-} reduced $\text{Ox Ag} + 3\text{Cs} \rightarrow \text{Cs}_3\text{Ag}$