

Chapter Test B**Chapter: Acids and Bases**

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. Which of the following is an oxyacid?
a. HCl
b. H_2O
c. H_2S
d. H_2SO_4
- _____ 2. Which acid is produced in the stomach?
a. HNO_3
b. CH_3COOH
c. H_2SO_4
d. HCl
- _____ 3. Which of the following is a strong base?
a. NH_3
b. aniline
c. NaOH
d. acetate ion
- _____ 4. In the reaction represented by the equation $HF(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + F^-(aq)$, a conjugate acid-base pair is
a. F^- and H_2O .
b. H_3O^+ and HF.
c. HF and F^- .
d. HF and H_2O .
- _____ 5. If H_2O in the reaction represented by the equation $H_2O + C_2H_3COOH \rightleftharpoons H_3O^+ + C_2H_3COO^-$ is considered to be a weaker base, then H_3O^+ is a
a. stronger acid.
b. stronger base.
c. weaker acid.
d. weaker base.
- _____ 6. Proton-transfer reactions favor production of the
a. stronger acid and stronger base.
b. weaker acid and weaker base.
c. stronger acid and weaker base.
d. weaker acid and stronger base.
- _____ 7. Aqueous solutions of most bases contain
a. hydroxide ions and cations.
b. hydroxide ions and anions.
c. hydrogen ions and anions.
d. hydrogen ions and cations.
- _____ 8. Acid strength increases with
a. increasing polarity and increasing bond strength.
b. increasing polarity and decreasing bond strength.
c. decreasing polarity and increasing bond strength.
d. decreasing polarity and decreasing bond strength.

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