

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks) $\text{pH} = 4.65$ so $[\text{H}_3\text{O}^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} \text{ M}$

$$\text{HBrO} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{BrO}^-$$

[I]	0.25		0	0
[C]	-2.239×10^{-5}		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		2.239×10^{-5}	2.239×10^{-5}

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{BrO}^-]}{[\text{HBrO}]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer $K_a = 2.0 \times 10^{-9}$

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4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $\text{pH} = 2.355$ $[\text{H}_3\text{O}^+] = \text{antilog}(-2.355) = 0.004416 \text{ M}$

$$\text{C}_6\text{H}_5\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{C}_6\text{H}_5\text{COO}^-$$

[I]	C_0		0	0
[C]	-0.004416		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		0.004416	0.004416

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_6\text{H}_5\text{COO}^-]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

Answer $C_0 = 0.30 \text{ M}$

$$C_0 = 0.299976 + 0.004416 = 0.30439 \text{ M}$$

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5. Find the value of Kb for the oxalate ion ($\text{C}_2\text{O}_4^{2-}$). (1 mark)

$$K_b(\text{C}_2\text{O}_4^{2-}) = \frac{K_w}{K_a(\text{HC}_2\text{O}_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

Answer $K_b = 1.6 \times 10^{-10}$

1

6. The value of Kb for the weak base methylamine (CH_3NH_2) is 4.4×10^{-4} . Calculate the value of Ka for the acid CH_3NH_3^+ . (1 mark)

$$K_a(\text{CH}_3\text{NH}_3^+) = \frac{K_w}{K_b(\text{CH}_3\text{NH}_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

Answer $K_a = 2.3 \times 10^{-11}$

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Basic Stoichiometry Phet Lab Answer Key

Wolfgang Guggemos



Basic Stoichiometry Phet Lab Answer Key :

The book delves into Basic Stoichiometry Phet Lab Answer Key . Basic Stoichiometry Phet Lab Answer Key is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Basic Stoichiometry Phet Lab Answer Key , encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Basic Stoichiometry Phet Lab Answer Key
- Chapter 2: Essential Elements of Basic Stoichiometry Phet Lab Answer Key
- Chapter 3: Basic Stoichiometry Phet Lab Answer Key in Everyday Life
- Chapter 4: Basic Stoichiometry Phet Lab Answer Key in Specific Contexts
- Chapter 5: Conclusion

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