

## Titration Calculation Practice Questions

1. Eggshells contain calcium carbonate. The following procedure was used to determine the percentage, by mass, of calcium carbonate in 1.13 g of eggshell.

**Step 1** The 1.13 g of eggshell were crushed and added to 100.0 mL of 0.300 mol L<sup>-1</sup> HCl solution. Excess HCl remained after the reaction was complete.



**Step 2** 20.0 mL samples of the excess HCl were titrated with 0.10 mol L<sup>-1</sup> Na<sub>2</sub>CO<sub>3</sub> solution.

- (a) Calculate the initial number of moles of HCl present before the reaction in Step 1.  
(b) The equation for the titration reaction in Step 2 is shown below:



In one titration 8.35 mL of Na<sub>2</sub>CO<sub>3</sub> was needed to neutralise the HCl in one 20.0 mL sample.

- (i) Calculate the number of moles of Na<sub>2</sub>CO<sub>3</sub> needed to neutralise the HCl in Step 2.  
(ii) Hence state the number of moles of HCl in the 20.0 mL sample titrated in Step 2.  
(iii) Hence calculate the total number of moles of excess HCl that remained after the reaction with the eggshell in Step 1.  
(c) Calculate the number of moles of HCl that reacted with the eggshell in Step 1.  
(d) Calculate the mass of calcium carbonate in the 1.13 g of eggshell.  
(e) Calculate the percentage mass of calcium carbonate in eggshell.

2. The following procedure was used to determine the concentration of hydrogen gas in a sample of air:

**Step 1** 1.0 × 10<sup>2</sup> L of air was bubbled through 0.0500 L of 0.300 mol L<sup>-1</sup> acidified K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> solution. An equation for the reaction that occurred is shown below:



Excess Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> remained in the solution after the reaction.

**Step 2** The excess Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> was titrated with 2.00 mol L<sup>-1</sup> Fe<sup>2+</sup> solution. An equation for the reaction that occurred is shown below:



A titre value of 17.75 mL was obtained.

- (a) Calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> present before the reaction with H<sub>2</sub> in Step 1.  
(b) Calculate the number of moles of Fe<sup>2+</sup> required to react with the Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> in Step 2.  
(c) Hence calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> left unreacted after Step 1.  
(d) Hence calculate the number of moles of Cr<sub>2</sub>O<sub>7</sub><sup>2-</sup> that reacted with H<sub>2</sub> in Step 1.  
(e) Calculate the number of moles of H<sub>2</sub> in the 1.0 × 10<sup>2</sup> L of air.  
(f) Calculate the concentration, in %w/v, of hydrogen gas in the air.

# Titration Questions And Answers

**AW Rasmussen**



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