

Inorganic Chemistry

Lesson 1: Inorganic Chemistry

Inorganic Chemistry - deals with **MATTER** its composition, properties and its reactions.

Matter - is anything that occupies space and has mass.

- All the objects in the universe, since they occupy space and have mass, are matter.
- The mass of an object is the quantity of matter that it contains.

Classification of Matter:

1. **PURE SUBSTANCES / SUBSTANCES** have fixed compositions and distinct properties.

- **ELEMENTS** - are substances that cannot be decomposed into simpler substances by chemical means. They are the basic substances out of which all matter is composed.
 - Metal
 - Nonmetal
 - Metalloid

COMPOUNDS - are the substances composed of two or more elements united chemically in definite proportions by mass. They can be decomposed by chemical means into simpler elements.

- Acid
- Base
- Salt

2. **MIXTURES** are composed of two or more kinds of matter where each substance retains its own identity and properties thus it can be separated by physical means. Compositions of mixtures can vary continuously.

→ **HOMOGENEOUS MIXTURES (1 phase / SOLUTIONS)** - have uniform composition throughout. Solutions are classified under this group, i.e. *vinegar, brine, syrup.*

→ **HETEROGENEOUS MIXTURES (2/ SUSPENSIONS)** - are substances whose composition, properties, and appearance differ throughout, i.e. sand in water, chocolate chip ice cream, fruit salad.

- ◆ **silverware** - a homogeneous mixture (alloy)

Properties of Matter - refer to the characteristics that enable one to distinguish one substance to another.

1. **Physical properties:** color, odor, taste, hardness, physical state, melting point, boiling point, electrical conductivity, malleability, density, solubility, specific gravity.
2. **Chemical properties:** rusting metal, burning metal, burning fuel, milk turning sour

Physical States Matter

1. **Solid** - the molecules that make up a solid are arranged in regular, repeating patterns. They are held firmly in place but can vibrate within a limited area.
2. **Liquid** - the molecules that make up a liquid flow easily around one another. They are kept from flying apart by attractive forces between them. Liquids assume the shape of their containers.
3. **Gas** - the molecules that make up a gas fly in all directions at great speeds. They are so far apart that attractive forces between them are insignificant.
4. **Plasma** - at the very high temperatures of stars, atoms lose their electrons. The mixture of electrons and nuclei that results is the plasma state of matter.

Changes of Matter

1. **Physical change** does not involve a change of one of matter into another; no change in the fixed composition of the substance.

Inorganic Photochemistry Lecture Notes

SJ Ball



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Table of Contents Inorganic Photochemistry Lecture Notes

1. Understanding the eBook Inorganic Photochemistry Lecture Notes
 - The Rise of Digital Reading Inorganic Photochemistry Lecture Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Inorganic Photochemistry Lecture Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Inorganic Photochemistry Lecture Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Inorganic Photochemistry Lecture Notes
 - Personalized Recommendations

- Inorganic Photochemistry Lecture Notes User Reviews and Ratings
- Inorganic Photochemistry Lecture Notes and Bestseller Lists
- 5. Accessing Inorganic Photochemistry Lecture Notes Free and Paid eBooks
 - Inorganic Photochemistry Lecture Notes Public Domain eBooks
 - Inorganic Photochemistry Lecture Notes eBook Subscription Services
 - Inorganic Photochemistry Lecture Notes Budget-Friendly Options
- 6. Navigating Inorganic Photochemistry Lecture Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Inorganic Photochemistry Lecture Notes Compatibility with Devices
 - Inorganic Photochemistry Lecture Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Inorganic Photochemistry Lecture Notes
 - Highlighting and Note-Taking Inorganic Photochemistry Lecture Notes
 - Interactive Elements Inorganic Photochemistry Lecture Notes
- 8. Staying Engaged with Inorganic Photochemistry Lecture Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Inorganic Photochemistry Lecture Notes
- 9. Balancing eBooks and Physical Books Inorganic Photochemistry Lecture Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Inorganic Photochemistry Lecture Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Inorganic Photochemistry Lecture Notes
 - Setting Reading Goals Inorganic Photochemistry Lecture Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Inorganic Photochemistry Lecture Notes
 - Fact-Checking eBook Content of Inorganic Photochemistry Lecture Notes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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