



<p>1. Which of the following best describes what causes water molecules to exhibit properties such as cohesive and adhesive behavior?</p>	<p>Water molecules are polar, having partially positive and partially negative ends.</p>
<p>2. Which statement correctly compares the pH of water to that of other substances?</p>	<p>The pH of water is neutral; it is higher than the pH of acids and lower than the pH of bases.</p>
<p>3. The slight negative charge at one end of one water molecule is attracted to the slight positive charge of another water molecule. What is this attraction called?</p>	<p>A hydrogen bond</p>
<p>4. Which of the following best explains how water is able to dissolve ionic substances?</p>	<p>The positive and negative charges of the water molecule pull on the negative and positive ends of the ionic molecule to break them apart.</p>
<p>5. During winter's freezing temperatures, some plants can remove water from their cells to reduce the amount of water that will freeze inside them. Why would plants need to do this?</p>	<p>The expansion of excess water as it turns to ice inside the cells would break the cells apart.</p>
<p>6. Which of the following best describes how the density of water differs from that of other substances?</p>	<p>The density of water depends on its temperature, unlike other substances.</p>

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