



1.	Which of the following best describes what causes water molecules to exhibit properties such as cohesive and adhesive behavior?	Water molecules are polar, having partially positive and partially negative ends.
2.	Which statement correctly compares the pH of water to that of other substances?	The pH of water is neutral; it is higher than the pH of acids and lower than the pH of bases.
3.	The slight negative charge at one end of one water molecule is attracted to the slight positive charge of another water molecule. What is this attraction called?	A hydrogen bond
4.	Which of the following best explains how water is able to dissolve ionic substances?	The positive and negative charges of the water molecule pull on the negative and positive ends of the ionic molecule to break them apart.
5.	During winter's freezing temperatures, some plants can remove water from their cells to reduce the amount of water that will freeze inside them. Why would plants need to do this?	The expansion of excess water as it turns to ice inside the cells would break the cells apart.
6.	Which of the following best describes how the density of water differs from that of other substances?	The density of water depends on its temperature, unlike other substances.

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