

INTRODUCTION

As an official testing centre of the TEF exam, we are providing this guide to help prepare for the examination.

This guide is intended to help candidates from understand the overall structure of the TEF Canada , as well as, practice some of the type of exercises that they will encounter during the exam. It is not intended to replace any official guides nor is it a guarantee that the examples will be in the exam.

More examples and suggestions are available at the official page of the CCI Paris Ile-de-France Education. We strongly encourage candidates to review them.



The TEF Canada has different types of evaluations, depending on the type of application.

TEF CANADA IMMIGRATION



Oral comprehension : 40 minutes (60 questions)



Written comprehension: 60 minutes (50 questions)



Oral expression: 15 minutes (2 topics to cover)



Written expression: 60 minutes (2 topics to cover)

TEF CANADA CITIZENSHIP



Oral comprehension: 40 minutes (60 questions)



Oral expression: 15 minutes (2 topics to cover)

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$$2 \text{Cu} + 2 \text{AgNO}_3 \rightarrow 2 \text{Cu(NO}_3)_2 + 2 \text{Ag}$$

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you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0

g Al_2O_3 $\times 1 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$ 26.982 g Al_2O_3 $\times 4 \text{ mol Al}_2\text{O}_3 = 107.728 \text{ g Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4.33 g of H_2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

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reaction

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