

F.E. Sem - 2 CBCGS SPA – MAY 17

Q.1.a.) Convert 0010 0100 0010 1101 from base 2 to decimal. Convert 134 from base 10 to hexadecimal. Write steps for conversion.

Ans:

Binary to decimal: The steps to be followed are:

1. Multiply the binary digits (bits) with powers of 2 according to their positional weight.
2. The position of the first bit (going from right to left) is 0 and it keeps on incrementing as you go towards left for every bit.
3. Given: $(0010\ 0100\ 0010\ 1101)_2 = (?)_{10}$
4. $0 \cdot 2^{15} + 0 \cdot 2^{14} + 1 \cdot 2^{13} + 0 \cdot 2^{12} + 0 \cdot 2^{11} + 1 \cdot 2^{10} + 0 \cdot 2^9 + 0 \cdot 2^8 + 0 \cdot 2^7 + 0 \cdot 2^6 + 1 \cdot 2^5 + 0 \cdot 2^4 + 1 \cdot 2^3 + 1 \cdot 2^2 + 0 \cdot 2^1 + 1 \cdot 2^0$
 $= 0 + 0 + 8192 + 0 + 0 + 1024 + 0 + 0 + 0 + 0 + 32 + 0 + 8 + 4 + 1$
 $= 9261$
 $(0010\ 0100\ 0010\ 1101)_2 = (9261)_{10}$

Decimal to hexadecimal: The steps to be followed are:

1. Divide the decimal number by 16. Treat the division as an integer division.
2. Write down the remainder (in hexadecimal).
3. Divide the result again by 16. Treat the division as an integer division.
4. Repeat step 2 and 3 until result is 0.
5. The hexadecimal value is digit sequence of the reminders from last to first.
6. Given: $(134)_{10} = (?)_{16}$

	16	134	
16	8	6	↑
	0	8	

$$(134)_{10} = (86)_{16}$$

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standard molar volume figure 9 18

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web this lab explores the relationships between pressure volume and temperature in an ideal gas the ideal gas law $pV = nRT$ is used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

8 2 relating pressure volume amount and temperature the - May 05 2023

web figure pageindex 10 since the number of moles in a given volume of gas varies with pressure and temperature changes chemists use standard temperature and pressure 273 15 k and 1 atm or 101 325 kpa to report properties of gases

lab report 5 docx pressure volume and temperature - Aug 28 2022

web pressure volume and temperature discovery of gas laws experiment 5 10 3 19 jared caviglia section 582 ian nicholson introduction the ideal gas law is a commonly used formula in chemistry it relates pressure temperature and volume this experiment seeks to prove those relationships by measuring the pressure of gas at various temperatures

pressure volume temperature lab ku school of engineering - Jun 25 2022

web the pressure volume temperature pvt fluids lab in ku s chemical petroleum engineering department is used to study the properties of fluids under a wide range of pressure and temperature

lab report pressure volume and temperature discovery - Jan 01 2023

web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

pressure temperature relationship in gases lab studocu - Nov 30 2022

web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

pressure volume temperature pvt laboratory department - Sep 09 2023

web aug 13 2021 pressure volume temperature pvt laboratory analysis of fluid properties such as api gravity viscosity surface and interfacial tension ph refractive index and vapor pressure is performed in the pvt laboratory pressure volume and temperature relationships of hydrocarbons are studied at the graduate level

che144 thermodynamics engineering lab report saturation vapor pressure - Jul 27 2022

web the properties of water at constant volume can be represented as a function of pressure and temperature as shown in the diagram below figure 1 the pressure temperature relationship at constant volume the saturation point of water is the condition at which a phase change occurs from liquid to vapour or vapour to liquid

6 3 relationships among pressure temperature volume and - Jul 07 2023

web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts