

VSA QUESTIONS (1 - MARK QUESTIONS)

1. Define the term 'rate of reaction'.
2. Mention the units of rate of reaction.
3. Express the rate of reaction in terms of $\text{Br}^- (\text{aq})$ as reactant and $\text{Br}_2 (\text{aq})$ as product for the reaction :



4. For a chemical reaction represented by $\text{R} \rightarrow \text{P}$ the rate of reaction is denoted by

$$\frac{-\Delta[\text{R}]}{\Delta t} \text{ or } \frac{+\Delta[\text{P}]}{\Delta t}$$

Why a positive sign (+) is placed before $\frac{\Delta[\text{P}]}{\Delta t}$ and negative sign (-)

before $\frac{\Delta[\text{R}]}{\Delta t}$?

5. Express the rate of reaction in terms of disappearance of hydrogen and appearance of ammonia in the given reaction.

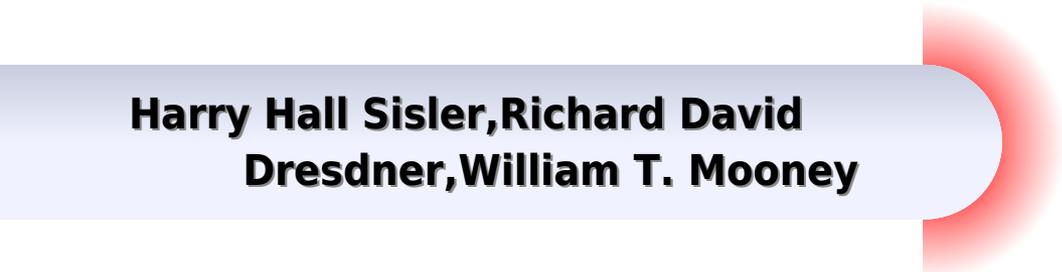


6. Why rate of reaction does not remain constant throughout?
7. Write the unit of first order rate constant of a gaseous reaction if the partial pressure of gaseous reactant is given in bar.
8. For a zero order reaction :

$\text{R} \rightarrow \text{P}$, the change in concentration of reactant w.r.t. time is shown by following graph.

Chemical Kinetics Questions And Answers

**Harry Hall Sisler, Richard David
Dresdner, William T. Mooney**



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Energetics, Kinetics, and Life George Tyler Miller, 1971

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Rates and Mechanisms of Chemical Reactions W. C. Gardiner (Jr.), William Cecil Gardiner, 1969 *Physical Chemistry: Physical and chemical kinetics* R. Stephen Berry, Stuart Alan Rice, John Ross, 1980 *Kaplan SAT: Chemistry* Claire Aldridge, Karl Lee, 2006-02 Includes 3 Full length practice tests Comprehensive review of all the tested material with practice quizzes for each chapter Effective strategies to maximize your score

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Introduction to Chemical Thermodynamics M. H. Everdell, 1965

Chemical Abstracts, 1925 **Process Chemistry for Water and Wastewater Treatment** Larry D. Benefield, Joseph F. Judkins, Barron L. Weand, 1982

The Alchemical Essence and the Chemical Element Muir, 1894 *Chemistry* Harry Hall Sisler, Richard David Dresdner, William T. Mooney, 1980 Vividly presents the fundamental of chemistry for science and engineering majors Offers the basics of the discipline the theory of chemistry and the role it plays in our understanding of the material world the descriptive chemistry of elements and compounds and the significance of chemical knowledge and its implications for social and environmental problems

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