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calculate the number of moles from each of the following a 6.95×10^{23} atoms of Cu b 7.12×10^{21} molecules of CO_2 c 3.33×10^{22} ions of Pb^{4+} d 7.00×10^{23} formula units of LiNO_3

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using the relationships found in the balanced chemical

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the quantity of heat for a process is represented by the letter q the sign of q for an endothermic process is positive because the system is gaining heat a chemical reaction or physical

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the quantities commonly encountered in chemical problems include the number of moles of a substance the number of atoms molecules or formula units of a substance and the mass in

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it is always a good idea to estimate the answer before you do the actual calculation in this case the mass given 35 00 g is less than the molar mass so the answer should be less than 1

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the atomic mass of an element expressed in grams is the mass of a mole of the element the mass of a mole of an element is the molar mass to calculate the molar mass of a compound

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the amount of a substance that contains 6.02×10^{23} representative particles of that substance the amount of a substance that contains 6.02×10^{23} representative particles of that

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