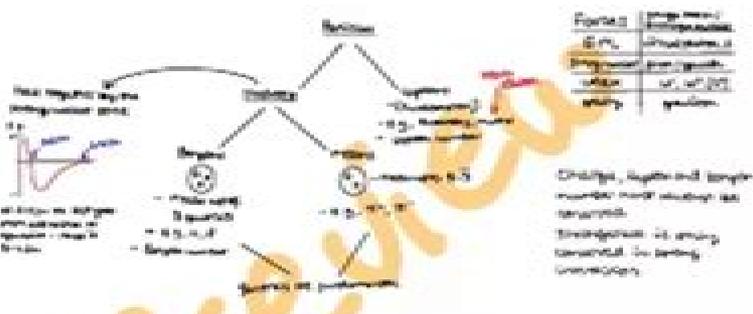


Particles and Quantum



Excitation: when a particle and its corresponding antiparticle collide and their rest energy is converted into photons.



Annihilation: when a particle turns into a particle and its corresponding antiparticle. The photon must have at least the same amount of energy as the rest energy of the particle.



Excitation

Excitation means they give electrons enough energy to climb their atomic shells.

- Alpha particles** - Helium nucleus, ^4_2He
- Highly ionising, so it is dangerous
- Relatively penetrating - stopped by a piece of paper or foil on their skin
- $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + ^4_2\text{He}$

Beta - fast moving electron $\rightarrow e^-$

- Medium penetrating and ionising ability - stopped by a few cm of aluminium
- Beta decay: $^{14}_6\text{C} \rightarrow ^{14}_7\text{N} + e^- + \bar{\nu}_e$ (antineutrino is almost undetectable)



Gamma - high energy EM ray or photon emitted from a high energy nucleus $\rightarrow \gamma$

- Highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons

- different relative atomic mass - different number of neutrons



Nuclear charge - $\frac{Q}{4\pi\epsilon_0} \left(\frac{1}{r^2} \right)$

Electron volt - charge of an electron when accelerated through a pot of one volt

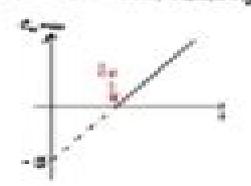
- $1.6 \times 10^{-19} \text{ J}$
- 1 eV



Photoelectric effect

Use photons of sufficient energy on electrons on the surface of the metal, liberating them.

- $E_{\text{photon}} = hf = \phi + E_{\text{kin}}$
- 1 - photon absorbed - hf (total energy)
- 2 - frequency
- 3 - work function - minimum energy to liberate electrons
- 4 - kinetic energy left over after ϕ is subtracted
- 5 - stopped potential - minimum potential needed for electrons to be liberated



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons E_{kin} is not linearly dependent on the wavelength λ but is linearly dependent on the frequency f of light. E_{kin} is zero for frequencies below f_0 and increases linearly for frequencies above f_0 .

Stopping potential - the pot. needed to stop any electrons from reaching the other side.

- $E_{\text{photon}} = hf$
- ϕ - charge of electron
- eV_s - stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have.

The wave nature of particles was proven by diffraction of a crystal lattice which produces Bragg's law of diffraction. This is because the electron is diffracting within the crystal lattice.



- $\lambda = \frac{h}{p}$
- $\lambda = \frac{h}{mv}$
- λ - de Broglie wavelength
- p - momentum / mass x velocity

Fluorescent tubes



These tubes contain mercury under a high voltage is applied. This excites the gas molecules to which some of the mercury atoms emit photons when they relax.

When the gas relaxes it emits photons when the electrons in the gas are excited. The photons are in the UV range. A phosphor coating in the tube absorbs the photons and emits visible light.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In a gas discharge tube, when an electron collides with a gas atom, it can transfer its energy.

Ionisation energy - is enough energy to pull an electron out of an atom, so will completely change the atom.



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Raffaela Di Napoli



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