

## Key Equations for Colligative Properties

Raoult's Law: Relationship between the Vapor Pressure of a Solution ( $P_{\text{solution}}$ ), the Mole Fraction of the Solvent ( $\chi_{\text{solvent}}$ ), and the Vapor Pressure of the Pure Solvent ( $P^{\circ}_{\text{solvent}}$ )

$$P_{\text{solution}} = \chi_{\text{solvent}} P^{\circ}_{\text{solvent}}$$

$$\chi_{\text{solvent}} = \frac{n_{\text{solvent}}}{n_{\text{solvent}} + n_{\text{solute}}}$$

The Vapor Pressure of a Solution Containing Two Volatile Components

$$P_A = \chi_A P^{\circ}_A \quad P_B = \chi_B P^{\circ}_B$$

$$P_{\text{total}} = P_A + P_B$$

Relationship between Osmotic Pressure ( $\Pi$ ), Molarity ( $M$ ), the Ideal Gas Constant ( $R$ ), and Temperature ( $T$ , in K).

$$\Pi = MRT$$

$$M = \frac{n_{\text{solute}}}{V_{\text{solution}}} \quad 0.08206 \text{ L}\cdot\text{atm} / \text{mol}\cdot\text{K}$$

The dependence of  $\Delta T_f$  and  $\Delta T_b$  on the Molality ( $m$ ), and Freezing Point Depression Constant ( $K_f$ ) and Boiling Point Elevation Constant ( $K_b$ ) respectively.

$$\Delta T_f = m \cdot K_f$$

$$\Delta T_b = m \cdot K_b$$

$$m = \frac{\text{mol (solute)}}{\text{kg (solvent)}}$$

# Colligative Properties Problems And Solutions

**D Keegan**



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