

AS Level Chemistry

11: Group 17

ALLEN HARRING
CAPITAL

- Cl_2 (g) : pale yellow gas
 Br_2 (l) : reddish brown liquid
 I_2 (s) : black solid / purple vapour
- colour intensity increases down Grp 17.

- 6 p block elements
 oxidation number: -1 or 0. (can go up to +7 bc of vacant d orbitals - except F)

Atomic Radius

Tendency change overwhelmed by 7 shielding, nuclear attraction remains same \downarrow radius, **atomic radius \uparrow**

MP/BP & Volatility

simple molecular structures \rightarrow weak D-D than molecules \rightarrow no. of electrons $\uparrow \rightarrow$ **boil pt \uparrow** \rightarrow **mp/bp \uparrow**
 \rightarrow diff. to separate molecules \rightarrow **volatility \downarrow**

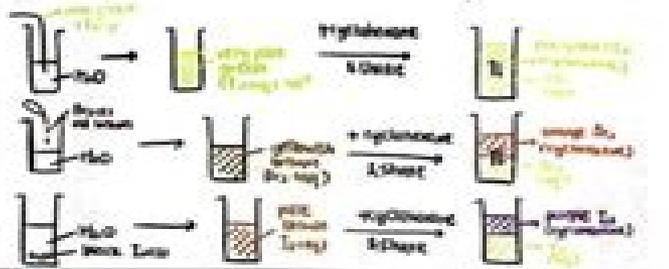
Solubility in water

oxidising power $\downarrow \rightarrow$ solubility \downarrow

- F: - oxides water vigorously even at -100K.
 $2\text{F}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 4\text{HF}(\text{aq}) + \text{O}_2(\text{g})$
- Cl: - disproportionates partially in water
 $\text{Cl}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCl}(\text{aq}) + \text{HOCl}(\text{aq})$
 - chlorides more decompose w/ strong oxidant
 $2\text{HClO}(\text{aq}) \rightarrow 2\text{HCl}(\text{aq}) + \text{O}_2(\text{g})$
- Br: - moderately soluble in water \rightarrow forms equilibrium with mixture of $\text{Br}_2(\text{aq})$ - HBrO formed negligible.
- I: - dissolves slightly in water to give pale yellow soln.
 - reacts in KI solⁿ to form I_3^- (brown solⁿ)
 $\text{I}_2(\text{s}) + \text{I}^-(\text{aq}) \rightarrow \text{I}_3^-(\text{aq})$

Solubility in Cyclohexane

- insoluble in water \rightarrow forms 2 separate layers
 \rightarrow halogens dissolved \rightarrow distinctly diff. colours



Bond Energy

atomic radii $\uparrow \rightarrow$ **less effective orbital overlap** \rightarrow less attraction for shared pair \rightarrow increasing bond length \rightarrow **bond energy \downarrow** from Cl to I.

Why F-F bond isn't longest:

smallest valence orbitals \rightarrow LP of F repel each other \rightarrow weakens F-F bond (like phosphorus w/ arsenic soln increases bc non bonding electrons have more space to spread out)

11.2 OXIDISING AGENT



- molecular size \uparrow , **oxidising power \downarrow**
- not readily to gain e^-** bc nuclear nuclear attraction
- less likely to be reduced
- oxidising power \downarrow

straight **A*** student's notes

A Level Chemistry Revision Notes

C. Boes



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