

Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid, $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$, with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The K_a for butanoic acid is 1.54×10^{-5} .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M HClO_4 after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M $\text{Ba}(\text{OH})_2$ solution have been added to 25.00 mL of 0.0400 M HNO_3 solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

Acid Base Titration Practice Problems With Answers

Xiang Xie



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Unveiling the Magic of Words: A Overview of "**Acid Base Titration Practice Problems With Answers**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Acid Base Titration Practice Problems With Answers**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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