

Atomic structure and the periodic table

Atomic structure



protons
neutrons
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

Equations, mixtures and chromatography

chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

products

word equation: methane + oxygen → carbon dioxide + water

symbol equation: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$



nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

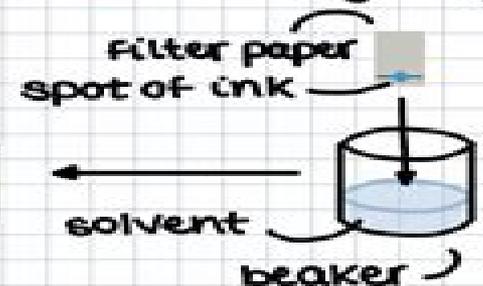
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

paper chromatography



Chapter 7 Notes Atomic Structure And Periodicity

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CHEMISTRY OF THE ENVIRONMENT Introduction to Modern Chemistry: Student Manual Eugene Meyer, 1979

Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg

Patricia L. Samuel, 1984 Gen Chem Irm Darrell D. Ebbing, 1996-06 *Mendeleev to Oganesson* Eric Scerri, Guillermo Restrepo, 2018-02-13 Since 1969 the international chemistry community has only held conferences on the topic of the Periodic Table three times and the 2012 conference in Cusco Peru was the first in almost a decade The conference was highly interdisciplinary featuring papers on geology physics mathematical and theoretical chemistry the history and philosophy of chemistry and chemical education from the most reputable Periodic Table scholars across the world Eric Scerri and Guillermo Restrepo have collected fifteen of the strongest papers presented at this conference from the most notable Periodic Table scholars The collected volume will contain pieces on chemistry philosophy of science applied mathematics and science education General Chemistry Darrell D. Ebbing, 1999 **Introduction to Chemistry** Arthur L. Williams, Harland D. Embree, Harold J. DeBey, 1973 *Instructor's Resource Guide to Accompany Chemistry & Chemical Reactivity* John C. Kotz, 1996

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