

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the the systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{FeSCN}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s}) + \text{H}^{+}(\text{aq})$$
4. In part 1, what is the conjugate acid and base? In part 2, what is the conjugate acid and base?
Conjugate acid: FeSCN^{2+} is the conjugate acid and Fe^{3+} is the conjugate base. In part 2, Fe^{3+} is the conjugate acid and FeSCN^{2+} is the conjugate base. In part 3, FeSCN^{2+} is the conjugate acid and $\text{Fe}(\text{OH})_2$ is the conjugate base.
5. What is the conjugate acid and base? In part 1, what is the conjugate acid and base?
Conjugate acid: FeSCN^{2+} is the conjugate acid and Fe^{3+} is the conjugate base. In part 2, Fe^{3+} is the conjugate acid and FeSCN^{2+} is the conjugate base. In part 3, FeSCN^{2+} is the conjugate acid and $\text{Fe}(\text{OH})_2$ is the conjugate base.
6. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s})$$
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s})$$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s})$$
9. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
$$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_2(\text{s})$$

Investigating Chemical Equilibrium Lab 12a Answer

Christian G. Meyer



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