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 $2 \text{ ch}_4 + 2 \text{o}_2 \rightarrow 2 \text{co}_2 + 2 \text{h}_2\text{o}$
 $2 \text{ na} + \text{cl}_2 \rightarrow 2 \text{nacl}$
 $3 \text{ al} + \text{o}_2 \rightarrow \text{al}_2\text{o}_3$
 $4 \text{ al} + 3 \text{o}_2 \rightarrow 2 \text{al}_2\text{o}_3$
 $4 \text{ n}_2 + \text{h}_2 \rightarrow 2 \text{nh}_3$

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 $\text{h}_2 + \text{o}_2 \rightarrow \text{h}_2\text{o}$
 at the molecular level the reaction would look something like this notice that there are two oxygen atoms on the

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web oct 13 2023 this means that you will need to balance the carbon atoms first 5 use a coefficient to balance the single carbon atom add a coefficient to the single carbon atom on the right of the equation to balance it with the 3 carbon atoms on the left of the equation
 $\text{c}_3\text{h}_8 + \text{o}_2 \rightarrow \text{h}_2\text{o} + 3 \text{co}_2$

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