

## Acid-Base Titration Practice Problems

- 1) Calculate the pH during the titration of 40.00 mL of 0.0500 M butanoic acid,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CO}_2\text{H}$ , with 0.1000 M NaOH solution after the following additions of the base: a) 0.00 mL; b) 15.00 mL; c) 19.00 mL; d) 20.00 mL; e) 25.00 mL. The  $K_a$  for butanoic acid is  $1.54 \times 10^{-5}$ .
- 2) Calculate the pH during the titration of 25.00 mL of 0.200 M KOH with 0.100 M  $\text{HClO}_4$  after the following additions of acid: a) 0.00 mL; b) 25.00 mL; c) 39.00 mL; d) 50.00 mL; e) 60.00 mL.
- 3) Calculate the pH when the following quantities of 0.0500 M  $\text{Ba}(\text{OH})_2$  solution have been added to 25.00 mL of 0.0400 M  $\text{HNO}_3$  solution: a) 0.00 mL; b) 5.00 mL; c) 10.00 mL; d) 20.00 mL.

# Acid Base Titration Problems With Answers

**James L. Flowers, Princeton Review  
(Firm)**



## **Acid Base Titration Problems With Answers:**

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**5 Steps to a 5 AP Chemistry, 2012-2013 Edition** Richard H. Langley, John Moore, 2011-06-10 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You ll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the Knowledge Step 5 Build Your Confidence Topics include Basics Reactions and Periodicity Stoichiometry Gases Thermodynamics Spectroscopy Light and Electrons Bonding Solids Liquids and Intermolecular Forces Solutions and Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry Organic Chemistry Experimental

**Analytical Chemistry, International Adaptation** Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2021-04-20 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students

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*Applied Physical Pharmacy, Third Edition* Mansoor Amiji, Thomas J. Cook, W. Cary Mobley, 2019-06-26 A complete practice oriented introduction to physical pharmacy Written to clearly and simply explain how drugs work this textbook explores the fundamental physicochemical attributes and processes important for understanding how a drug is transformed into a usable product that is administered to a patient to reach its pharmacological target and then exists the body Applied Physical Pharmacy Third Edition begins with a review of the key biopharmaceutics concepts of drug liberation absorption distribution metabolism and excretion These concepts and others set the framework for the subsequent chapters that describe physicochemical properties and process related to the fate of the drug Other physical pharmacy topics important to drug formulation are discussed in the chapters that follow which describe dispersal systems interfacial phenomena and rheology The textbook concludes with an overview of the principles of kinetics that are important for understanding the rates at which many of the processes discussed in previous chapters occur Chapters in this Third Edition retain the acclaimed learning aids of previous editions including Learning Objectives Practice Problems Key Points and Clinical Questions In order to be of greater value to the pharmacy student more clinical questions have been added and many tables have been updated with more current products and excipients Understanding Chemistry Fred M. Dewey, 1994 *Principles of Modern Chemistry*

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