

# Acids and Bases Worksheet 1

1. Finish the following reactions when acids are added to water.



2. What are the conjugate bases of these acids?

original acid	conjugate base
$\text{HNO}_3$	
$\text{H}_2\text{O}$	
$\text{H}_3\text{O}^+$	
$\text{H}_2\text{SO}_4$	
$\text{HBr}$	
$\text{HCO}_3^-$	

3. What are the conjugate acids of these bases?

original base	conjugate acid
$\text{OH}^-$	
$\text{H}_2\text{O}$	
$\text{HCO}_3^-$	
$\text{SO}_4^{2-}$	
$\text{ClO}_4^-$	

4. Which of the following represent conjugate acid-base pairs?



5. Calculate the  $[\text{H}^+]$  in a solution in which  $[\text{OH}^-] = 2.0 \times 10^{-2} \text{ M}$ . Is this solution acidic, neutral, or basic?

6. Calculate the  $[\text{OH}^-]$  in a solution in which  $[\text{H}^+] = 3.99 \times 10^{-5} \text{ M}$ . Is this solution acidic, neutral, or basic?

# Acids And Bases Study Guide Answer Key

**ML Yell**



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## **Table of Contents Acids And Bases Study Guide Answer Key**

1. Understanding the eBook Acids And Bases Study Guide Answer Key
  - The Rise of Digital Reading Acids And Bases Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids And Bases Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Study Guide Answer Key
  - Personalized Recommendations
  - Acids And Bases Study Guide Answer Key User Reviews and Ratings
  - Acids And Bases Study Guide Answer Key and Bestseller Lists
5. Accessing Acids And Bases Study Guide Answer Key Free and Paid eBooks
  - Acids And Bases Study Guide Answer Key Public Domain eBooks
  - Acids And Bases Study Guide Answer Key eBook Subscription Services
  - Acids And Bases Study Guide Answer Key Budget-Friendly Options

6. Navigating Acids And Bases Study Guide Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Acids And Bases Study Guide Answer Key Compatibility with Devices
  - Acids And Bases Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases Study Guide Answer Key
  - Highlighting and Note-Taking Acids And Bases Study Guide Answer Key
  - Interactive Elements Acids And Bases Study Guide Answer Key
8. Staying Engaged with Acids And Bases Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acids And Bases Study Guide Answer Key
9. Balancing eBooks and Physical Books Acids And Bases Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acids And Bases Study Guide Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases Study Guide Answer Key
  - Setting Reading Goals Acids And Bases Study Guide Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases Study Guide Answer Key
  - Fact-Checking eBook Content of Acids And Bases Study Guide Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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