

A weak base is titrated by a form of the conjugate acid. Consider the titration of methylamine, CH_3NH_2 , with hydrochloric acid, HCl . The methyl group of an amine is not acidic. The lone pair on the nitrogen atom is basic. The methyl group of an acid is the proton that you want to add.



Example:

When 1 mole of a weak base reacts with 1 mole of a strong acid, the products are 1 mole of a weak acid and 1 mole of a strong base.

Write the balanced chemical equation for the reaction between methylamine and hydrochloric acid.



Write the balanced chemical equation for the reaction between methylamine and acetic acid.

Write the balanced chemical equation for the reaction between methylamine and acetic acid.



Write the following equations. Be sure to use correct chemical formulas and states.

- Write the balanced chemical equation for the reaction between methylamine and hydrochloric acid. Be sure to use correct chemical formulas and states. $\text{CH}_3\text{NH}_2 + \text{HCl} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{Cl}^-$
- Write the balanced chemical equation for the reaction between methylamine and acetic acid. $\text{CH}_3\text{NH}_2 + \text{CH}_3\text{COOH} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{CH}_3\text{COO}^-$
- Write the balanced chemical equation for the reaction between methylamine and acetic acid. $\text{CH}_3\text{NH}_2 + \text{CH}_3\text{COOH} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{CH}_3\text{COO}^-$
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Judd E. Hollander



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