

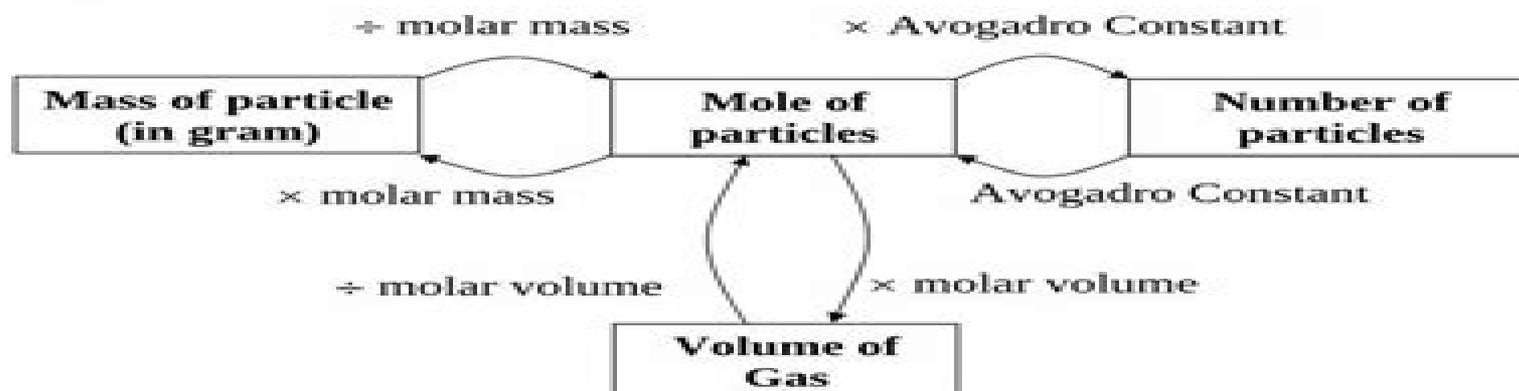
Short Notes: Form 4 Chemistry

Chemical Formulae and Equation

Calculation

<p>For Solid, liquid or gas</p> $\text{number of mole} = \frac{\text{mass of substance}}{\text{molar mass}}$ <p>Molar mass = RAM/RMM/RFM in gram</p>	<p>For gas (only)</p> $\text{number of mole} = \frac{\text{volume of gas}}{\text{molar volume}}$ <p>Molar volume = 24dm³ at room temperature Molar volume = 22.4dm³ at s.t.p.</p>
<p>For Solution</p> $\text{number of mole} = \frac{MV}{1000}$ <p>M = molarity V = Volume of solution in cm³</p>	<p>For quantity of particle(atom,molecule,ion)</p> $\text{number of mole} = \frac{\text{quantity of particle}}{6.02 \times 10^{23}}$

Summary



One School Short Notes Form 4 Chemistry

Yan Bai



One School Short Notes Form 4 Chemistry:

Decoding **One School Short Notes Form 4 Chemistry**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**One School Short Notes Form 4 Chemistry**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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