

## CHAPTER 2

2.1 Two possible versions can be developed:

```

IF x ≥ 10 THEN
  DO
    x = x - 5
    IF x < 50 EXIT
  END DO
ELSE
  IF x < 5 THEN
    x = 5
  ELSE
    x = 7.5
  END IF
ENDIF

```

```

IF x ≥ 10 THEN
  DO
    x = x - 5
    IF x < 50 EXIT
  END DO
ELSEIF x < 5
  x = 5
ELSE
  x = 7.5
ENDIF

```

2.2

```

DO
  i = i + 1
  IF x > 50 EXIT
  x = x + 5
  IF x > 5 THEN
    y = x
  ELSE
    y = 0
  ENDIF
  z = x + y
ENDDO

```

2.3 Students could implement the subprogram in any number of languages. The following VBA program is one example. It should be noted that the availability of complex variables in languages such as Fortran 90 would allow this subroutine to be made even more concise. However, we did not exploit this feature, in order to make the code more compatible with languages that do not support complex variables.

```

Option Explicit

Sub Rootfind()
  Dim ier As Integer
  Dim a As Double, b As Double, c As Double
  Dim r1 As Double, i1 As Double, r2 As Double, i2 As Double
  a = 7: b = 6: c = 2
  Call Roots(a, b, c, ier, r1, i1, r2, i2)
  If ier = 0 Then
    MsgBox "No roots"
  ElseIf ier = 1 Then
    MsgBox "single root=" & r1
  ElseIf ier = 2 Then
    MsgBox "real roots = " & r1 & ", " & r2
  ElseIf ier = 3 Then
    MsgBox "complex roots = " & r1 & ", " & i1 & "i" & ", " & r2 & ", " & i2 & "i"
  End If
End Sub

Sub Roots(a, b, c, ier, r1, i1, r2, i2)
  Dim d As Double
  r1 = 0: r2 = 0: i1 = 0: i2 = 0

```

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**Tobias Bleicker**



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical  $\times 100$  what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno<sub>3</sub> h<sub>2</sub>so<sub>4</sub> in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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**lab report 15 nitration of bromobenzene studocu** - Aug 31 2023

lab report 15 nitration of bromobenzene brynn abbott amp katie muerer bea0006 uah ch 336 02 september 21 2022 abstract in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

nitration of bromobenzene the purpose of this lab was to - Jul 30 2023

the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no<sub>2</sub><sup>+</sup> which acts as the electrophilic species

**nitration of bromobenzene lab report docx lab report 10** - Oct 21 2022

nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0ml of nitric acid and 4.0ml of sulfuric acid and will then react with 4.5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

**experiment 11 electrophilic aromatic substitution nitration** - Feb 10 2022

experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

**nitration of bromobenzene chem 154 gwu studocu** - Mar 26 2023

lab writeup the purpose of this lab was to observe where substituents are added on benzene ring way of electrophilic aromatic substitution the electrophilic

**5 nitration lab docx lab 5 nitration of bromobenzene** - Aug 19 2022

chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially

produced chemicals by adding different functional groups into an aromatic ring

lab report nitration of bromobenzene ipl org - Feb 22 2023

lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group