

CHEM 103 Lab Exam 3 (100 out of 100) Questions and Verified Answers (Portage Learning)

Scientific Calculator

Question 1

5 / 10 pts

Answer the following two questions:

1. Why was the hydrate heated with the Fisher burner in part 2 of lab 3?
2. For what purpose was SEM carried out in lab 3?

1. The hydrate was heated with the Fisher burner in lab 3 to cause loss of water.
2. The SEM was carried out in lab 3 to detect the elemental composition.

2. The SEM was carried out in lab 3 to detect the elemental composition. (We did not use it for imaging)

Question 2

10 / 10 pts

Scientific Calculator

Answer the following question.

A 1.1000 gram hydrate sample chosen from $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$, $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$, $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ and $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ was heated and found to lose 0.6920 gram of H_2O . (1) Show the calculation of the % H_2O in the unknown hydrate sample. (2) Show the calculation of the % H_2O in each of the hydrate compounds and identify the unknown hydrate from the list.

Atomic weights: H = 1.008, O = 16.00. MWs: $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ = 286.15, $\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$ = 241.43, $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ = 203.301 and $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ = 244.462

% H_2O in unknown = $(0.6920 / 1.1000) \times 100 = 62.91\%$

% H_2O in $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$ = $180.16 / 286.15 \times 100 = 62.96\%$

unknown is **$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$**

Question 3

10 / 10 pts

Chem 110 Lab Questions And Answers

**Forest Products Laboratory
(U.S.), Raymond N. Miller, W. H.
Swanson**

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