

Covalent Bonding Worksheet

Covalent bonding occurs when two or more **NON_METALS** share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.) H_2 hydrogen is diatomic: $H + H \longrightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{H}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{H}} \longrightarrow H-H \quad \text{Single Bond}$
2.) F_2 fluorine is diatomic: $F + F \longrightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{\underset{\cdot\cdot}{F}}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{\underset{\cdot\cdot}{F}}} \longrightarrow$
3.) O_2 oxygen is diatomic: $O + O \longrightarrow \quad \quad \quad \longrightarrow O=O \quad \text{Double Bond}$
4.) N_2 nitrogen is diatomic. Is this a triple bond? $N + N$
5.) BF_3 you need 3 fluorine atoms here: $B + F$
6.) Ammonia NH_3 hint: how many hydrogen atoms are needed? $N + H$
7.) Carbon dioxide CO_2 $O + C + O$
8.) Methane CH_4 careful here 4 hydrogen atoms needed $C + H$
9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!! $H + O + H$
10.) SO_2 hint: one pair of electrons from sulfur must be slit up for this one to work. $O + S + O$

Chapter 8 Covalent Bonding Worksheet Answers

M Woodhall



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