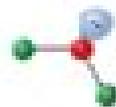


Steric # = Total e- pairs on central atom	Electron Distribution	# Bond Pairs	# Lone Pairs	General AXE Formula	Molecular Shape	e.g.	Lewis Structure	3-D Shape (Sketch)	Bond Angle(s)	Polar Molecule? P or NP (4.5)
2	Linear	1	0	AX	Linear	H <sub>2</sub>	H-H		180°	NP
2	Linear	2	0	AX <sub>2</sub>	Linear	BeCl <sub>2</sub>			180°	NP
3	Trigonal Planar	3	0	AX <sub>3</sub>	Trigonal Planar	BCl <sub>3</sub>			120°	NP
		2	1	AX <sub>2</sub> E	Bent	SO <sub>2</sub>			<120°	P ≠ EN >0.4
4	Tetrahedral	4	0	AX <sub>4</sub>	Tetrahedral	CCl <sub>4</sub>			109.5°	NP
		3	1	AX <sub>3</sub> E	Trigonal Pyramidal	NCl <sub>3</sub>			107°	P ≠ EN >0.4
		2	2	AX <sub>2</sub> E <sub>2</sub>	Bent	OCl <sub>2</sub> H <sub>2</sub> O			104.5°	P ≠ EN >0.4

# Dot Vsepr Lab Class Set

**M Walker**



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