

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A <u>great deal of energy is produced</u> when an ionic compound is formed. |

4 1 Review Reinforcement Answer Key

Jessica J Manson



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Table of Contents 4 1 Review Reinforcement Answer Key

1. Understanding the eBook 4 1 Review Reinforcement Answer Key
 - The Rise of Digital Reading 4 1 Review Reinforcement Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 4 1 Review Reinforcement Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 4 1 Review Reinforcement Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 4 1 Review Reinforcement Answer Key
 - Personalized Recommendations
 - 4 1 Review Reinforcement Answer Key User Reviews and Ratings
 - 4 1 Review Reinforcement Answer Key and Bestseller Lists
5. Accessing 4 1 Review Reinforcement Answer Key Free and Paid eBooks
 - 4 1 Review Reinforcement Answer Key Public Domain eBooks
 - 4 1 Review Reinforcement Answer Key eBook Subscription Services
 - 4 1 Review Reinforcement Answer Key Budget-Friendly Options

6. Navigating 4 1 Review Reinforcement Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 4 1 Review Reinforcement Answer Key Compatibility with Devices
 - 4 1 Review Reinforcement Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 4 1 Review Reinforcement Answer Key
 - Highlighting and Note-Taking 4 1 Review Reinforcement Answer Key
 - Interactive Elements 4 1 Review Reinforcement Answer Key
8. Staying Engaged with 4 1 Review Reinforcement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 4 1 Review Reinforcement Answer Key
9. Balancing eBooks and Physical Books 4 1 Review Reinforcement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 4 1 Review Reinforcement Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 4 1 Review Reinforcement Answer Key
 - Setting Reading Goals 4 1 Review Reinforcement Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 4 1 Review Reinforcement Answer Key
 - Fact-Checking eBook Content of 4 1 Review Reinforcement Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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