



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - H_2O_2
 - CO_2

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- K_2SO_4
- C_2H_6
- H_2O
- H_2SO_4
- CaCl_2
- $\text{C}_2\text{H}_5\text{OH}$
- H_2O_2
- CO_2

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. CH_4
- Show covalent bonding in carbon dioxide. CO_2
- Show covalent bonding in ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for each of the following molecules. Show all covalent bonding explicitly.
 - H_2O
 - CO_2
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_2$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_3$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_4$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_5$
 - $\text{C}_2\text{H}_6\text{N}_2\text{O}_6$



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital do you combine to form the σ bond in ethane and acetylene?
- Which orbital and sub-orbital form the π bond in acetylene?

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