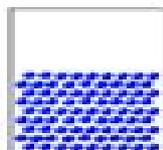


Last Name _____ First Name _____ Due Date _____ Period _____

Kinetic Molecular Theory Worksheet

Use the following terms to fill in the blanks. Some may be used more than once. Hopefully all will be included but you never know. If you seem to need an additional term (oops!) just do your best

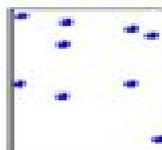
Gas	Solid	Liquid	Phase	Compress
Empty Space	Kinetic Energy	Bonds	Structure	Collisions
Anomaly	Random Motion	Temperature	Absolute Zero	density
Fluids	Compressible	Incompressible		



A



B



C

A	SOLID
B	LIQUID
C	GAS

The boxes above represent matter in the 3 common phases. Write the name of the phase next to the appropriate letter.

- The word kinetic comes from a Greek word that means "to move." The kinetic molecular theory is based upon the assumption that atoms are in constant **Random Motion**.
- State the names of the phases in order of Temperature, Kinetic Energy, density and bond strength between atoms.

Lowest Temperatures	In between Temperatures	Highest Temperatures
Solid	Liquid	Gas
Lowest Kinetic Energy	In between Kinetic Energy	Highest Kinetic Energy
Solid	Liquid	Gas
Lowest Density		Highest Density
Gas		
High Strength Bonds	Weak Bonds	No Bonds
Solid	Liquid	Gas

- Kinetic Energy** is directly related to temperature. As the temperature increases so does the **Kinetic Energy**.
- Compared to Solids and Liquids gases tend to have a lower **density**. This can be explained by the fact that there is so much **Empty Space** between particles in a gas.
- A small pressure will easily **compress** a gas. If this happens its **density** will increase. The space between the atoms will (INCREASE or **DECREASE**).
- Even a large pressure will not **compress** a solid because the atoms are so closely packed.
- Solids have strong **Bonds**, liquids have weaker **Bonds** and gases have no **Bonds** holding them together.
- A solid like ice or diamond has a highly ordered arrangement of bonds between atoms which is called its **Structure**. A liquid or gas does not have this.

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M Mark



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Table of Contents Kinetic Molecular Theory Worksheet Answers

1. Understanding the eBook Kinetic Molecular Theory Worksheet Answers
 - The Rise of Digital Reading Kinetic Molecular Theory Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Kinetic Molecular Theory Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kinetic Molecular Theory Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetic Molecular Theory Worksheet Answers
 - Personalized Recommendations
 - Kinetic Molecular Theory Worksheet Answers User Reviews and Ratings
 - Kinetic Molecular Theory Worksheet Answers and Bestseller Lists

5. Accessing Kinetic Molecular Theory Worksheet Answers Free and Paid eBooks
 - Kinetic Molecular Theory Worksheet Answers Public Domain eBooks
 - Kinetic Molecular Theory Worksheet Answers eBook Subscription Services
 - Kinetic Molecular Theory Worksheet Answers Budget-Friendly Options
6. Navigating Kinetic Molecular Theory Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Kinetic Molecular Theory Worksheet Answers Compatibility with Devices
 - Kinetic Molecular Theory Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kinetic Molecular Theory Worksheet Answers
 - Highlighting and Note-Taking Kinetic Molecular Theory Worksheet Answers
 - Interactive Elements Kinetic Molecular Theory Worksheet Answers
8. Staying Engaged with Kinetic Molecular Theory Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kinetic Molecular Theory Worksheet Answers
9. Balancing eBooks and Physical Books Kinetic Molecular Theory Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kinetic Molecular Theory Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Kinetic Molecular Theory Worksheet Answers
 - Setting Reading Goals Kinetic Molecular Theory Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kinetic Molecular Theory Worksheet Answers
 - Fact-Checking eBook Content of Kinetic Molecular Theory Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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