

Synthesis of Alum

1. Second Quiz (Density/Alum)
2. Collect lab-report (Density) and Pre-Lab (Alum)
3. Lecture on Today's Experiment
4. Demonstrate how to perform gravity filtration, suction filtration and Bunsen burner

Lecture:

Goal: synthesize alum from Aluminum can

- Learn to perform laboratory technique of purification by filtration
- Calculating theoretical yield and percent yield

Chemical formula for Alum : $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$

Alum – use in medical, cooking, water treatment etc,.....

Chemical RXN – Handout



→ step 3-4 : Dissolve metal in excess strong base KOH to form a complex ion and H_2 gas (bubbles)

→ use gravity filtration to filter the solution while it's still hot, keep the filtrate (What is the filtrate)



→ step 8: Neutralize excess $-\text{OH}$ in the filtrate with H_2SO_4 → precipitation will form



→ Step 9: more acid is added to dissolve $\text{Al}(\text{OH})_3$ solid to form a soluble $\text{KAl}(\text{SO}_4)_2$, this is not really an Hydrated Alum



→ this can be filter using suction filtration and dry

→ Student must wait after 48 hours to weigh their sample and must be done prior the next class meeting (penalty will be applied)

→ Prefer to the class schedule to check when the lab will open, you should come 1 hour after the lab start and must ask for permission (WEAR GOGGLES!!!!)

Alum From Aluminum Lab Answers

K Morrison



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