

# AS Level Chemistry

## 11: Group 17

ALLEN HARRING  
CAPITAL

- $\text{Cl}_2$  (g) : pale yellow gas  
 $\text{Br}_2$  (l) : reddish brown liquid  
 $\text{I}_2$  (s) : black solid / purple vapour
- } colour intensity increases down Grp 17.

- 6 p block elements  
 oxidation number: -1 or 0. (can go up to +7 bc of vacant d orbitals - except F)

### Atomic Radius

Tendency change overwhelmed by shielding, nuclear attraction remains same  $\downarrow$  radius, **atomic radius  $\uparrow$**

### MP/BP & Volatility

simple molecular structures  $\rightarrow$  weak D-D than molecules  $\rightarrow$  no. of electrons  $\uparrow \rightarrow$  **boil pt  $\uparrow$**   $\rightarrow$  **mp/bp  $\uparrow$**   
 $\rightarrow$  diff. to separate molecules  $\rightarrow$  **volatility  $\downarrow$**

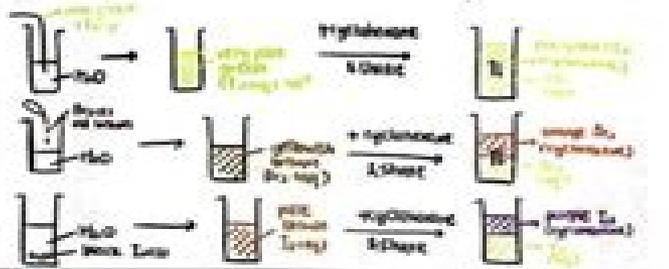
### Solubility in water

oxidising power  $\downarrow \rightarrow$  solubility  $\downarrow$

- F: - oxides water vigorously even at -100K.  
 $2\text{F}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 4\text{HF}(\text{aq}) + \text{O}_2(\text{g})$
- Cl: - disproportionates partially in water  
 $\text{Cl}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCl}(\text{aq}) + \text{HOCl}(\text{aq})$   
 - chlorides are decomposed w/ strong oxidant  
 $2\text{HClO}(\text{aq}) \rightarrow 2\text{HCl}(\text{aq}) + \text{O}_2(\text{g})$
- Br: - moderately soluble in water  $\rightarrow$  forms equilibrium with mixture of  $\text{Br}_2(\text{aq})$  -  $\text{HBrO}$  formed negligible.
- I: - dissolves slightly in water to give pale yellow soln.  
 - reacts in KI sol<sup>n</sup> to form  $\text{I}_3^-$  (brown sol<sup>n</sup>)  
 $\text{I}_2(\text{s}) + \text{I}^-(\text{aq}) \rightarrow \text{I}_3^-(\text{aq})$

### Solubility in Cyclohexane

- insoluble in water  $\rightarrow$  forms 2 separate layers  
 $\rightarrow$  halogens dissolved  $\rightarrow$  distinctly diff. colours



### Bond Energy

atomic radii  $\uparrow \rightarrow$  **less effective orbital overlap**  $\rightarrow$  less attraction for shared pair  $\rightarrow$  increasing bond length  $\rightarrow$  **bond energy  $\downarrow$**  from Cl to I.

Why F-F bond isn't longest:

smallest valence orbitals  $\rightarrow$  LP of F repel each other  $\rightarrow$  weakens F-F bond (like phosphorus w/ arsenic soln increases bc non bonding electrons have more space to spread out)

### 11.2 OXIDISING AGENT

strongest ox



weakest ox

- molecular size  $\uparrow$ , **oxidising power  $\downarrow$**
- not readily to gain  $\text{e}^-$**  bc nuclear nuclear attraction
- less likely to be reduced
- oxidising power  $\downarrow$

straight **A\*** student's notes

# A Level Chemistry Revision Notes

**JL Elias**



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