

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups and subshells. The groups are numbered 1 through 18. The subshells are labeled as follows:

- 1s subshell:** Group 1 (Hydrogen, Alkali Metals)
- 2s subshell:** Groups 2 and 3 (Alkaline Earth Metals, Transition Metals)
- 3s subshell:** Groups 11 and 12 (Copper Group, Zinc Group)
- 4s subshell:** Groups 1 and 2 (Alkali Metals, Alkaline Earth Metals)
- 5s subshell:** Groups 1 and 2 (Alkali Metals, Alkaline Earth Metals)
- 6s subshell:** Groups 1 and 2 (Alkali Metals, Alkaline Earth Metals)
- 7s subshell:** Groups 1 and 2 (Alkali Metals, Alkaline Earth Metals)



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For a neutral atom, $Z = +$ the number of electrons. The atomic number of an element is the number of protons in the nucleus. The atomic weight of an element is the sum of the atomic weights of the isotopes of that element. The atomic weight of an element is the sum of the atomic weights of the isotopes of that element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an atomic orbital. The principal quantum number (n) determines the energy and size of the orbital. The azimuthal quantum number (l) determines the shape of the orbital. The magnetic quantum number (m_l) determines the orientation of the orbital. The spin quantum number (m_s) determines the spin of the electron in the orbital.

NUCLEAR CHEMISTRY

Nuclear reactions involve the nucleus of an atom. The nucleus is made of protons and neutrons. The mass number (A) is the sum of the number of protons and neutrons. The atomic number (Z) is the number of protons. The number of neutrons is $A - Z$.

MANY-ELECTRON ATOMS

In many-electron atoms, the energy of an orbital depends on both the principal quantum number (n) and the azimuthal quantum number (l). The energy levels are split into subshells. The order of filling of orbitals is given by the Aufbau Principle.

n	l	ml	orbital
1	0	0	1s
2	0	0	2s
2	1	0	2p _z
2	1	±1	2p _x , 2p _y
3	0	0	3s
3	1	0	3p _z
3	1	±1	3p _x , 3p _y
3	2	0	3d _{z²}
3	2	±1	3d _{xy} , 3d _{xz} , 3d _{yz}
3	2	±2	3d _{x²-y²}

Atomic configuration: The arrangement of electrons in an atom is described by the atomic configuration. The order of filling of orbitals is given by the Aufbau Principle. The Pauli Exclusion Principle states that no two electrons in an atom can have the same set of quantum numbers.



Detailed periodic table showing element symbols and names. The table is color-coded by groups and periods.

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