

ANSWERS TO CONVERSION FACTOR PROBLEMS

Conversion factors can be used to convert units or to convert between equivalent ways of expressing a quantity. The quantity in the problem is multiplied by one or more "conversion factors," in which the numerator is equal to the denominator. Since the numerator and denominator of the conversion factor are equal, multiplying by the conversion factor is like multiplying by 1 and thus does not change the value of the original quantity. Use the table of English to Metric equivalents as needed. All answers should be in *significant figures*!

A) Problems with a single conversion factor.

1. How many cm are in 18.9 inches?

$$18.9 \text{ in} \times \frac{2.54 \text{ cm}}{1 \text{ in}} = 47.5 \text{ cm}$$

2. How many grams are in 0.143 ounces?

$$0.143 \text{ oz} \times \frac{1 \text{ gram}}{0.03527 \text{ oz}} = 4.05 \text{ grams}$$

B) Problems with two or more conversion factors.

1. How many mL are in 0.037 quarts?

$$0.037 \text{ qt} \times \frac{1 \text{ L}}{1.057 \text{ qt}} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 35 \text{ mL}$$

2. How many grams are in 0.397 pounds (lbs)?

$$0.397 \text{ lb} \times \frac{1 \text{ kg}}{2.20 \text{ lb}} \times \frac{1000 \text{ g}}{1 \text{ kg}} = 180 \text{ g} = 1.80 \times 10^3 \text{ g}$$

3. How many micrograms are in 6.8×10^{-7} ounces?

$$6.8 \times 10^{-7} \text{ oz} \times \frac{1 \text{ g}}{0.0353 \text{ g}} \times \frac{1000 \text{ mg}}{1 \text{ g}} \times \frac{1000 \text{ } \mu\text{g}}{1 \text{ mg}} = 19 \text{ } \mu\text{g}$$

C) Using density as a conversion factor. The density of Al is 2.70 g/cm^3 .

Remember, $1 \text{ mL} = 1 \text{ cm}^3$.

1. What is the volume of 0.810 grams of Al?

$$0.810 \text{ g} \times \frac{1 \text{ cm}^3}{2.70 \text{ g}} = 0.300 \text{ cm}^3 \text{ (or } 0.300 \text{ mL)}$$

2. Find the mass in grams of 0.327 liters of Al.

$$0.327 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} \times \frac{2.70 \text{ g}}{1 \text{ cm}^3} = 883 \text{ g}$$

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**Russo Steve Silver Michael, Steve
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the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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