

Chapter 9: Chemical Reactions

- _____ 1. Which of the following could be an indication of a chemical reaction of a compound or element?
- a. It is a solid at room temperature.
 - b. It melts at 631^oC.
 - c. It expands upon freezing.
 - d. It burns in an atmosphere of chlorine.
- _____ 2. In this unbalanced chemical reaction, which components are the reactants?
- $$C_7H_{16} + O_2 \rightarrow CO_2 + H_2O$$
- a. C₇H₁₆ and O₂
 - b. C₇H₁₆ and CO₂
 - c. C₇H₁₆ and H₂O
 - d. CO₂ and H₂O
- _____ 3. When Ca(OH)₂ reacts with H₂SO₄, what are the products.
- a. Ca(OH)₂ and H₂SO₄
 - b. CaH₂ and H₂SO₄
 - c. CaSO₄ and H₂O
 - d. CaO and H₂SO₄
- Hint: This is a double-replacement reaction, positive ions switch places.
- _____ 4. What are the products of the reaction between barium hydroxide and hydrochloric acid?
- a. Barium chloride and water
 - b. Barium hydroxide and water
 - c. Barium chloride and barium hydride
 - d. Chloric acid and barium hydroxide
- Hint: Write out the skeleton equation. This is a double-replacement reaction.
- _____ 5. This symbol separates the reactants from products and means "to yield".
- a. +
 - b. (s)
 - c. \rightarrow
 - d. (g)
- _____ 6. In a chemical equation, the number written in front of the chemical formula is the
- a. Reactant
 - b. Subscript
 - c. Equation
 - d. Coefficient
- _____ 7. When there are exactly the same number and type of atoms on both sides of a chemical equation, the equation is _____.
- a. Balanced
 - b. Ionic
 - c. Synthetic
 - d. Decomposed
- _____ 8. The chemical equation $CaCO_3 \rightarrow CaO + CO_2$ is classified as a
- a. Synthesis reaction
 - b. Decomposition reaction
 - c. Single-replacement reaction
 - d. Double-replacement reaction
- _____ 9. In the combustion reaction of propane (C₃H₈) with oxygen gas, what are the products?
- a. C₃H₈ and O₂
 - b. C₃H₈ and CO₂
 - c. O₂ and H₂O
 - d. H₂O and CO₂
- _____ 10. The major physical indicator(s) of a combustion reaction is

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Mary Wuerth



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