

# GROWTH MINDSET

MISTAKES  
HELP ME  
LEARN

FEEDBACK IS  
VALUABLE

I IMPROVE  
WITH  
PRACTICE

IS THIS  
MY BEST  
WORK?

I WON'T  
GIVE UP



# FIXED MINDSET

I WANT TO  
AVOID MAKING  
MISTAKES

I'LL NEVER BE  
THAT SMART

I KNOW  
BEST

I  
GIVE UP

THIS IS  
GOOD  
ENOUGH



# Quick Start Self Help Mindset

**Peipei Pang**



## **Quick Start Self Help Mindset:**

Immerse yourself in heartwarming tales of love and emotion with its touching creation, Experience Love's Journey in **Quick Start Self Help Mindset** . This emotionally charged ebook, available for download in a PDF format ( \*), is a celebration of love in all its forms. Download now and let the warmth of these stories envelop your heart.

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web oxidation reduction redox reactions oxidation  $\text{Cu} + \text{Cu}^{2+} + 2\text{e}^-$  reduction  $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$  l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+$  is

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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{H}_2\text{SO}_3 + 2\text{H}^+ \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$  what will be the stoichiometric coefficient of  $\text{HSO}_4^-$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}^{2+} + \text{Fe}^{3+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{2+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}^+ + \text{H}_2\text{O}_2 + 2\text{Au}^{3+}$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^- + 3\text{Au}^{3+}$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_2 + 2\text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+}$  9  $\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} \rightarrow \text{TeO}_4 + 11\text{PbO}_2 + \text{IO}_2 + \text{Pb}_2\text{O}_3$  10  $\text{AsO}_4^{3-} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$  and  $\text{H}_2\text{O}$  in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + 2\text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} \rightarrow 2\text{HNO} + 2\text{C}_6\text{H}_8\text{O}_6 + 2\text{H}_2\text{O}$  choose 1 answer c  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{HNO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{HNO}$  b  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{H}_2\text{O}$  c  $\text{HNO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  d  $\text{HNO}_2$  and  $\text{C}_6\text{H}_6\text{O}_6$

