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Boston, Massachusetts
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Prentice Hall Earth Science Laboratory Manual Paperback

**Harold Connolly, Cyrena A.
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 b 4.25 g of NH_3 in 0.500 l of solution the concentration of NH_3 in household ammonia

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$$\text{molality} = \frac{\text{moles solute}}{\text{kilograms solvent}}$$
 mathematical manipulation of molality is the

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 molality $0.825 \text{ g} \times 141.9579 \text{ g/mol} \times 0.00581158 \text{ mol} \times 0.00581158 \text{ mol} \times 0.4500 \text{ kg} \times 0.0129 \text{ m}^3$
 mole fraction n_2 / n_{total}
 $0.825 \text{ g} \times 141.9579 \text{ g/mol} \times 0.00581158 \text{ mol} \times 2 \times 0.4500 \text{ g} \times 18.015 \text{ g/mol} \times 24.97918401 \text{ mol}$

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$$\text{begin array l} \text{ mass of a component} \frac{\text{mass of component in solution}}{\text{total mass of solution}} \times 100 \text{ end array}$$

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