

# Buffers Keep the Balance

## Post-Lab Questions

1. Compare the measured pH value for the model carbonate blood buffer to
  - a. the expected pH of an ideal carbonic acid-bicarbonate buffer, and
  - b. the actual pH of the carbonic acid-bicarbonate buffer system present in the blood.
2. Based on the pH comparisons in Question #1, which solution, the model carbonate blood buffer or an actual blood buffer, is more likely to contain a greater proportion of the carbonic acid component of the buffer compared to the bicarbonate component? Explain your reasoning.
3. What are the effects of adding even a small amount of HCl or NaOH on the pH value of the control solution (water)? Compare this to the effect of adding HCl or NaOH to the model carbonate blood buffer.
4. Which phosphate buffer solution in Part B corresponds to the composition of an ideal buffer solution? Compare its measured pH value with the calculated pH of the ideal buffer solution.

# Buffers Keep The Balance Lab Answers

**Yijin Wang**



## **Buffers Keep The Balance Lab Answers:**

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## The Enigmatic Realm of **Buffers Keep The Balance Lab Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Buffers Keep The Balance Lab Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of those that partake in its reading experience.

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