

Answer KEY

Unit 3 Outline/ Study Guide - Electron Configuration and Periodic Trends

Essential Skills/ State Standards:

1. Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, halogens, alkali metals, alkaline earth metals, and transition metals.
3. Students know how to use the periodic table to identify the following trends:
 ionization energy, electronegativity, and the relative sizes of ions and atoms
 (this includes knowing the definitions of ionization energy, electronegativity, cation, and anion.)

Study Guide - (use your notes and worksheets for a more complete review)

1. Identify the following things on the periodic table:

-metals	-alkali metals
-metalloids (semi-metals)	-alkaline earth metals
-transition metals	-halogens
-nonmetals	-noble gases
-Periods	-Groups/families

Identify groups with the following numbers of valence electrons:
 1, 2, 3, 4, 5, 6, 7, 8

Key:
 Metals (red)
 Metalloids (purple)
 Nonmetals (blue)

Group 1: Alkali metals
 Group 2: Alkaline earth metals
 Groups 3-10: Transition metals
 Group 17: Halogens
 Group 18: Noble gases

2. Why do the elements in the groups (columns) behave so similar to each other even if they have great differences in their number of protons? **same number of valence electrons and similar chemical properties.**
3. What is the connection between the following terms: **octet rule, valence electrons, and ions?**
In order to achieve the octet rule, atoms gain or lose their valence electrons. They become charged ions when this occurs.
4. a. List the elements that are exceptions to the octet rule b) Why don't these atoms achieve the octet rule?
 a. Small atoms are happy with 2 electrons (ex: H, He, Li, Be, B)
 b. because the 1s energy level only holds 2 electrons

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Lingsheng Yao



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Table of Contents Atoms Periodic Table Review Answer Key Lbrsfs

1. Understanding the eBook Atoms Periodic Table Review Answer Key Lbrsfs
 - The Rise of Digital Reading Atoms Periodic Table Review Answer Key Lbrsfs
 - Advantages of eBooks Over Traditional Books
2. Identifying Atoms Periodic Table Review Answer Key Lbrsfs
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atoms Periodic Table Review Answer Key Lbrsfs
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atoms Periodic Table Review Answer Key Lbrsfs
 - Personalized Recommendations
 - Atoms Periodic Table Review Answer Key Lbrsfs User Reviews and Ratings
 - Atoms Periodic Table Review Answer Key Lbrsfs and Bestseller Lists
5. Accessing Atoms Periodic Table Review Answer Key Lbrsfs Free and Paid eBooks

- Atoms Periodic Table Review Answer Key Lbrsfs Public Domain eBooks
 - Atoms Periodic Table Review Answer Key Lbrsfs eBook Subscription Services
 - Atoms Periodic Table Review Answer Key Lbrsfs Budget-Friendly Options
6. Navigating Atoms Periodic Table Review Answer Key Lbrsfs eBook Formats
- ePub, PDF, MOBI, and More
 - Atoms Periodic Table Review Answer Key Lbrsfs Compatibility with Devices
 - Atoms Periodic Table Review Answer Key Lbrsfs Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Atoms Periodic Table Review Answer Key Lbrsfs
 - Highlighting and Note-Taking Atoms Periodic Table Review Answer Key Lbrsfs
 - Interactive Elements Atoms Periodic Table Review Answer Key Lbrsfs
8. Staying Engaged with Atoms Periodic Table Review Answer Key Lbrsfs
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atoms Periodic Table Review Answer Key Lbrsfs
9. Balancing eBooks and Physical Books Atoms Periodic Table Review Answer Key Lbrsfs
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atoms Periodic Table Review Answer Key Lbrsfs
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Atoms Periodic Table Review Answer Key Lbrsfs
- Setting Reading Goals Atoms Periodic Table Review Answer Key Lbrsfs
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atoms Periodic Table Review Answer Key Lbrsfs
- Fact-Checking eBook Content of Atoms Periodic Table Review Answer Key Lbrsfs
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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