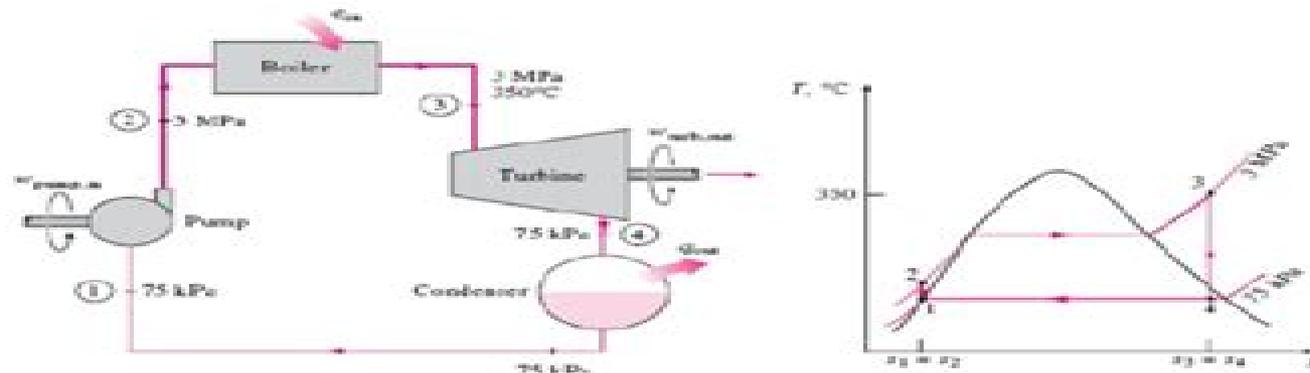


1. Consider a steam power plant operating on the simple ideal Rankine cycle. Steam enters the turbine at 3 MPa and 350°C and is condensed in the condenser at a pressure of 75 kPa. Determine the thermal efficiency of this cycle.

Solution A steam power plant operating on the simple ideal Rankine cycle is considered. The thermal efficiency of the cycle is to be determined.

Assumptions 1 Steady operating conditions exist. 2 Kinetic and potential energy changes are negligible.

Analysis The schematic of the power plant and the T - s diagram of the cycle are shown in Fig. 10–3. We note that the power plant operates on the ideal Rankine cycle. Therefore, the pump and the turbine are isentropic; there are no pressure drops in the boiler and condenser, and steam leaves the condenser and enters the pump as saturated liquid at the condenser pressure.



It is also interesting to note the thermal efficiency of a Carnot cycle operating between the same temperature limits

$$\eta_{th,Carnot} = 1 - \frac{T_{min}}{T_{max}} = 1 - \frac{(91.76 + 273) \text{ K}}{(350 + 273) \text{ K}} = 0.415$$

The difference between the two efficiencies is due to the large external irreversibility in Rankine cycle caused by the large temperature difference between steam and combustion gases in the furnace.

- 2–3. A steam power plant operates on the cycle shown in Fig. 10–5. If the isentropic efficiency of the turbine is 87 percent and the isentropic efficiency of the pump is 85 percent, determine (a) the thermal efficiency of the cycle and (b) the net power output of the plant for a mass flow rate of 15 kg/s. (a) The thermal efficiency of a cycle is the ratio of the net work output to the heat input, and it is determined as follows:

Pump work input:

$$\begin{aligned} W_{pump,in} &= \frac{W_{s,pump,in}}{\eta_p} = \frac{v_1(P_2 - P_1)}{\eta_p} \\ &= \frac{(0.001009 \text{ m}^3/\text{kg})[(16,000 - 9) \text{ kPa}]}{0.85} \left(\frac{1 \text{ kJ}}{1 \text{ kPa} \cdot \text{m}^3} \right) \\ &= 19.0 \text{ kJ/kg} \end{aligned}$$

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Table of Contents Thermodynamics Problem And Solution Mutinyore

1. Understanding the eBook Thermodynamics Problem And Solution Mutinyore
 - The Rise of Digital Reading Thermodynamics Problem And Solution Mutinyore
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Problem And Solution Mutinyore
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Problem And Solution Mutinyore
 - User-Friendly Interface

4. Exploring eBook Recommendations from Thermodynamics Problem And Solution Mutinyore
 - Personalized Recommendations
 - Thermodynamics Problem And Solution Mutinyore User Reviews and Ratings
 - Thermodynamics Problem And Solution Mutinyore and Bestseller Lists
5. Accessing Thermodynamics Problem And Solution Mutinyore Free and Paid eBooks
 - Thermodynamics Problem And Solution Mutinyore Public Domain eBooks
 - Thermodynamics Problem And Solution Mutinyore eBook Subscription Services
 - Thermodynamics Problem And Solution Mutinyore Budget-Friendly Options
6. Navigating Thermodynamics Problem And Solution Mutinyore eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermodynamics Problem And Solution Mutinyore Compatibility with Devices
 - Thermodynamics Problem And Solution Mutinyore Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Problem And Solution Mutinyore
 - Highlighting and Note-Taking Thermodynamics Problem And Solution Mutinyore
 - Interactive Elements Thermodynamics Problem And Solution Mutinyore
8. Staying Engaged with Thermodynamics Problem And Solution Mutinyore
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Problem And Solution Mutinyore
9. Balancing eBooks and Physical Books Thermodynamics Problem And Solution Mutinyore
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Problem And Solution Mutinyore
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermodynamics Problem And Solution Mutinyore
 - Setting Reading Goals Thermodynamics Problem And Solution Mutinyore
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Thermodynamics Problem And Solution Mutinyore
 - Fact-Checking eBook Content of Thermodynamics Problem And Solution Mutinyore
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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