

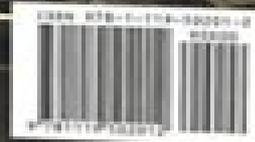
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web oxidation reduction redox reactions oxidation $\text{Cu} + \text{S} + \text{Cu} + 2\text{A} + \text{Q} + 2\text{E} + \text{Reduction} + \text{MnO} + 4\text{A} + \text{A} + \text{Q} + 4\text{H} + \text{A} + \text{Q} + 3\text{E} + \text{MnO} + 2\text{S} + 2\text{H} + 2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{S} + \text{MnO} + 4\text{A} + \text{A} + \text{Q}$ is

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^-$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$ the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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