

Name _____

Multiple Choice (10 points each)

Identify the letter of the choice that best completes the statement or answers the question.

- E 1. For a certain process at 27°C, $\Delta G = +210.6$ kJ and $\Delta H = -168.2$ kJ. What is the entropy change for this process at this temperature? Express your answer in the form, $\Delta S =$ _____ J/K.
- a. +628.3 J/K
 b. -141.3 J/K
 c. +141.3 J/K
 d. 1.26×10^3 J/K
 e. -1.26×10^3 J/K

Answer: $\Delta G = \Delta H - T\Delta S$ $T = 27^\circ\text{C} + 273 = 300\text{K}$

$$\Delta S = \frac{\Delta H - \Delta G}{T} = \frac{(-168.2\text{kJ}) - (+210.6\text{kJ})}{300\text{K}} = -\frac{1.26\text{kJ}}{\text{K}}$$

$$\left(\frac{-1.26\text{kJ}}{\text{K}}\right)\left(\frac{10^3\text{J}}{\text{kJ}}\right) = \boxed{\frac{-1.26 \times 10^3\text{J}}{\text{K}}}$$

- A 2. Which response includes all of the following substances that have $\Delta H_f^\circ = 0$, and no other substances?

- I. HCl(g)
 II. Na(s)
 III. HCl(aq)
 IV. F₂(g)

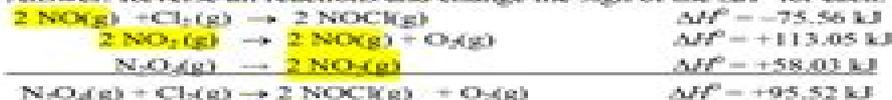
- a. II and IV
 b. II
 c. I and II
 d. I, II, III, and IV
 e. I, II, and IV

Answer: All elements in their natural standard state have a $\Delta H_f^\circ = 0$.

- A 3. From the following data at 25°C,

Calculate ΔH° at 25°C for the reaction below.

- a. +95.52 kJ
 b. +299 kJ
 c. -186.8 kJ
 d. -85.52 kJ
 e. -246.65 kJ

Answer: Reverse all reactions and change the sign of the ΔH° for each.

Thermodynamics Problems And Answers

Colin Bodsworth, A. S. Appleton



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Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy, 2020-03-23 *Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map* is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20-30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests

Chemical Engineering License Problems and Solutions Dilip K. Das, Rajaram K. Prabhudesai, 2003-09-18 This is a review book for people planning to take the PE exam in Chemical Engineering Prepared specifically for the exam used in all 50 states It features 188 new PE problems with detailed step by step solutions The book covers all topics on the exam and includes easy to use tables charts and formulas It is an ideal desk Companion to DAS's *Chemical Engineer License Review* It includes sixteen chapters and a short PE sample exam as well as complete references and an index Chapters include the following topical areas material and energy balances fluid dynamics heat transfer evaporation distillation absorption leaching liq liq extraction psychrometry and humidification drying filtration thermodynamics chemical kinetics process control mass transfer and plant safety The ideal study guide this book brings all elements of professional problem solving together in one BIG BOOK Ideal desk reference Answers hundreds of the most frequently asked questions The first truly practical no nonsense problems and solution book for the difficult PE exam Full step by step solutions are included

Solutions Manual for Quanta, Matter and Change Peter Atkins, Julio dePaula, Ron Friedman, 2008-12-15

Schaum's Outline of Theory and Problems of Thermodynamics for Engineers Merle C. Potter, Craig W. Somerton, 1995 This package develops the analysis of charge

carrying systems leading to an understanding of Maxwell's equations. Students can experiment with both advanced graphing and numerical techniques. Systems requirements are 80386/80486 PC or compatibles, Windows 3.1 or higher, 3.5 disk drive, 4 MB of RAM and 4 MB of disk space.

US Solutions Manual to Accompany Elements of Physical Chemistry 7e David Smith, 2017-09-28. The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book. The manual provides helpful comments and friendly advice to aid understanding. It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions.

The Thermodynamics of Electrical Phenomena in Metals, and A Condensed Collection of Thermodynamic Formulas Percy Williams Bridgman, 1961

Commonly Asked Questions in Thermodynamics Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-10. Have you ever had a question that keeps persisting and for which you cannot find a clear answer? Is the question seemingly so simple that the problem is glossed over in most resources or skipped entirely? CRC Press Taylor and Francis is pleased to introduce Commonly Asked Questions in Thermodynamics, the first in a new series of books that address the questions that frequently arise in today's major scientific and technical disciplines. Designed for a wide audience from students and researchers to practicing professionals in related areas, the books are organized in a user-friendly Question Answer format. Presented questions become increasingly specific throughout the book with clear and concise answers as well as illustrations, diagrams and tables are incorporated wherever helpful. Thermodynamics is a core discipline associated with the theoretical principles and practical applications underlying almost every area of science from nanoscale biochemical engineering to astrophysics. Highlighting chemical thermodynamics in particular, this book is written in an easy to understand style and provides a wealth of fundamental information, simple illustrations and extensive references for further research and collection of specific data. Designed for an audience that ranges from undergraduate students to scientists and engineers at the forefront of research, this indispensable guide presents clear explanations for topics with wide applicability. It reflects the fact that very often the most common questions are also the most profound.

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Physical Chemistry Student Solutions Manual Charles Trapp, Peter Atkins, Marshall Cady, 2006-08-11. Change 21

Classical and Quantum Thermal Physics R. Prasad, 2016-11-02. Discusses the interactions of heat energy and matter.

Thermal Physics and Statistical Mechanics S. K. Roy, 2001. This book emphasises the development of problem solving skills in undergraduate science and engineering students. The book provides more than 350 solved examples with complete step by step solutions as well as around 100 practice problems with answers. Also explains the basic theory, principles, equations and formulae for a quick understanding and review. Can serve both as a useful text and companion book to

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McGraw-Hill Education 500 Review Questions for the MCAT: General Chemistry Mary Millhollon, Richard H. Langley, 2015-07-10 500 ways to pass the General Chemistry section of the new MCAT Intensive practice detailed answer explanations the best way to sharpen skills and prepare for the exam In anticipation of the fully revised 2015 MCAT 500 Review Questions for the MCAT General Chemistry has been updated to comprehensively cover the chemistry portion of the Chemical and Physical Foundations of Biological Systems section This book provides the problem solving practice you need to take the exam with confidence 500 questions organized by subject Follows the new MCAT format Complete explanations to every question given in the answer key *Problems in Thermodynamics and Statistical Physics* Peter T.

Landsberg, 2014-07-16 Well respected and widely used this volume presents problems and full solutions related to a wide range of topics in thermodynamics statistical physics and statistical mechanics The text is intended for instructors undergraduates and graduate students of mathematics physics chemistry and engineering Twenty eight chapters each prepared by an expert proceed from simpler to more difficult subjects Similarly the early chapters are easier than the later ones making the book ideal for independent study Subjects begin with the laws of thermodynamics and statistical theory of information and of ensembles advancing to the ideal classical gases of polyatomic molecules non electrolyte liquids and solutions and surfaces Subsequent chapters explore imperfect classical and quantum gas phase transitions cooperative phenomena Green function methods the plasma transport in gases and metals Nyquist's theorem and its generalizations stochastic methods and many other topics

New Technical Books New York Public Library, 1915 **McGraw-Hill's 500 MCAT General Chemistry Questions to Know by Test Day** John T. Moore, Richard H. Langley, 2012-06-19 A wealth of problem solving practice in the format that you want This book is the ideal way to sharpen skills and prepare for this MCAT topic Get the problem solving practice for general chemistry you need with McGraw Hill's 500 MCAT General Chemistry Questions to Know by Test Day Organized for easy reference and intensive practice the questions cover all essential topics and the answer key includes detailed explanations for each question Inside you'll find 500 MCAT general chemistry questions organized by subject Detailed solutions to every problem given in the answer key Expert coverage for topics covered by the MCAT *Practical Thermodynamics* Forrest E. Cardullo, 1911 *Problems in Applied*

Thermodynamics Colin Bodsworth, A. S. Appleton, 1965
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