

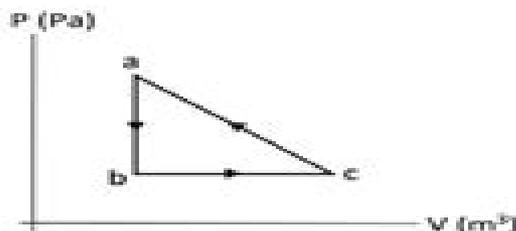
First Law of Thermodynamics

Grade

Name _____ Date _____

Use the following P-V diagram to answer questions about this thermodynamic cycle consisting of three processes labeled ab , bc , and ca .

$P_a = 3.00 \times 10^5 \text{ Pa}$, $P_b = 1.00 \times 10^5 \text{ Pa}$, $V_a = 3 \times 10^{-3} \text{ m}^3$, $V_c = 9 \times 10^{-3} \text{ m}^3$.



a) Identify the type of thermodynamic processes for each part of the cycle (ab , bc , ca) as:

- | | | | |
|--------|----------------|------------------|----------------|
| ab : | i) adiabatic | ii) isobaric | iii) isochoric |
| | iv) isothermal | v) none of these | |
| bc : | i) adiabatic | ii) isobaric | iii) isochoric |
| | iv) isothermal | v) none of these | |
| ca : | i) adiabatic | ii) isobaric | iii) isochoric |
| | iv) isothermal | v) none of these | |

b) For the process ab , determine (i) the work W_{ab} done by the gas and (ii) the change in internal energy of the system for a heat loss of 900 J. [Hints: for (i) recall that W is the area under the curve; for (ii) use the 1st Law.]

c) For the process bc , determine (i) the work W_{bc} done by the gas and (ii) the change in internal energy of the system for a heat gain of 1250 J.

d) For the process ca , determine (i) the work W_{ca} done by the gas and (ii) the change in internal energy of the system for a heat gain of 150 J.

e) Determine the total work W_{tot} done by the gas in the cycle. Compare this to the area of the triangle. Note that both are negative. Also note that the total change in internal energy is zero, as it must be when the system returns to the same state.

Thermodynamic Questions And Answers

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Table of Contents Thermodynamic Questions And Answers

1. Understanding the eBook Thermodynamic Questions And Answers
 - The Rise of Digital Reading Thermodynamic Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamic Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamic Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamic Questions And Answers
 - Personalized Recommendations
 - Thermodynamic Questions And Answers User Reviews and Ratings
 - Thermodynamic Questions And Answers and Bestseller Lists
5. Accessing Thermodynamic Questions And Answers Free and Paid eBooks

- Thermodynamic Questions And Answers Public Domain eBooks
 - Thermodynamic Questions And Answers eBook Subscription Services
 - Thermodynamic Questions And Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Thermodynamic Questions And Answers Compatibility with Devices
 - Thermodynamic Questions And Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamic Questions And Answers
 - Highlighting and Note-Taking Thermodynamic Questions And Answers
 - Interactive Elements Thermodynamic Questions And Answers
 8. Staying Engaged with Thermodynamic Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamic Questions And Answers
 9. Balancing eBooks and Physical Books Thermodynamic Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamic Questions And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermodynamic Questions And Answers
 - Setting Reading Goals Thermodynamic Questions And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermodynamic Questions And Answers
 - Fact-Checking eBook Content of Thermodynamic Questions And Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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